

Bonding Unit Test Review

Use your notes to answer the following questions.

Name _____ Period _____

1. What are valence electrons? **These are the electrons that are the farthest from the nucleus.**
2. The elements in a group have _____ **The same** _____ number of valence electrons.
3. Tell the number of valence electrons in each group:
 - a. Group 1 **1 valence electron**
 - b. Group 2 **2 valence electrons**
 - c. Group 13 **3 valence electrons**
 - d. Group 14 **4 valence electrons**
 - e. Group 15 **5 valence electrons**
 - f. Group 16 **6 valence electrons**
 - g. Group 17 **7 valence electrons**
 - h. Group 18 **8 valence electrons**
 - i. Groups 3-12 **# of valence electrons varies**
4. Cations lose valence electrons and have a _____ **positive** _____ ionic charge.
5. Anions gain valence electrons and have a _____ **negative** _____ ionic charge.
6. How do the electrons behave in a covalent bond?
The valence electrons are shared between atoms.
7. How do the electrons behave in an ionic bond?
The valence electrons are transferred. Atoms will either gain or lose electrons.
8. How do the electrons behave in a metallic bond?
Electrons flow freely between atoms.
9. Which type of bond dissolves easily in water?
Ionic Bonds
10. Which type of bond has a high melting point? Low melting point?
Ionic bonds and Metallic bonds have high melting points. Covalent bonds have a low melting point.
11. Which type of bond has the most energy stored in it?
Covalent Bond
12. Which type of bond conduct electricity as a solution? As a solid?
Ionic bonds conduct electricity as a solution. Metallic bonds conduct electricity as a solid.

Ionic Bonding Unit Test Review Answers

Kathleen Armour

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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