

Name Solutions

Date _____ Period _____

Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
Li	1 ve	transfers $1e^-$ to F	Li^+
F	7 ve	accepts $1e^-$ from Li	F^-

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $Li + Cl \Rightarrow LiCl$ $Li + \cdot\ddot{Cl}: \rightarrow [Li]^+ [\ddot{Cl}:]^{-1}$	Li	1 ve	donates $1e^-$	Li^+
	Cl	1 ve	accepts $1e^-$	Cl^{-1}
2) $Ca + O \Rightarrow CaO$ $Ca + \cdot\ddot{O}: \rightarrow [Ca]^{+2} [\ddot{O}:]^{-2}$	Ca	2 ve	donates $2e^-$	Ca^{+2}
	O	6 ve	accepts $2e^-$	O^{-2}
3) $Be + F \Rightarrow BeF_2$ $Be + \cdot\ddot{F}: \rightarrow [Be]^{+2} [\ddot{F}:]^{-1}$	Be	2 ve	donates $2e^-$ to each F	Be^{+2}
	F	7 ve	each F accepts $1e^-$ each	F^{-1}
4) $Mg + S \Rightarrow MgS$ $Mg + \cdot\ddot{S}: \rightarrow [Mg]^{+2} [\ddot{S}:]^{-2}$	Mg	2 ve	donates $2e^-$ to S	Mg^{+2}
	S	6 ve	accepts $2e^-$	S^{-2}
5) $K + F \Rightarrow KF$ $K + \cdot\ddot{F}: \rightarrow [K]^+ [\ddot{F}:]^{-1}$	K	1 ve	donates $1e^-$	K^{+1}
	F	7 ve	accepts $1e^-$	F^{-1}

Ionic Bonds Worksheet Answers Reading Notetaking

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Table of Contents Ionic Bonds Worksheet Answers Reading Notetaking

1. Understanding the eBook Ionic Bonds Worksheet Answers Reading Notetaking
 - The Rise of Digital Reading Ionic Bonds Worksheet Answers Reading Notetaking
 - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Bonds Worksheet Answers Reading Notetaking
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic Bonds Worksheet Answers Reading Notetaking
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Bonds Worksheet Answers Reading Notetaking
 - Personalized Recommendations
 - Ionic Bonds Worksheet Answers Reading Notetaking User Reviews and Ratings
 - Ionic Bonds Worksheet Answers Reading Notetaking and Bestseller Lists
5. Accessing Ionic Bonds Worksheet Answers Reading Notetaking Free and Paid eBooks
 - Ionic Bonds Worksheet Answers Reading Notetaking Public Domain eBooks
 - Ionic Bonds Worksheet Answers Reading Notetaking eBook Subscription Services
 - Ionic Bonds Worksheet Answers Reading Notetaking Budget-Friendly Options
6. Navigating Ionic Bonds Worksheet Answers Reading Notetaking eBook Formats

- ePub, PDF, MOBI, and More
 - Ionic Bonds Worksheet Answers Reading Notetaking Compatibility with Devices
 - Ionic Bonds Worksheet Answers Reading Notetaking Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic Bonds Worksheet Answers Reading Notetaking
 - Highlighting and Note-Taking Ionic Bonds Worksheet Answers Reading Notetaking
 - Interactive Elements Ionic Bonds Worksheet Answers Reading Notetaking
 8. Staying Engaged with Ionic Bonds Worksheet Answers Reading Notetaking
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ionic Bonds Worksheet Answers Reading Notetaking
 9. Balancing eBooks and Physical Books Ionic Bonds Worksheet Answers Reading Notetaking
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic Bonds Worksheet Answers Reading Notetaking
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Ionic Bonds Worksheet Answers Reading Notetaking
 - Setting Reading Goals Ionic Bonds Worksheet Answers Reading Notetaking
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Ionic Bonds Worksheet Answers Reading Notetaking
 - Fact-Checking eBook Content of Ionic Bonds Worksheet Answers Reading Notetaking
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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