

**Assessment 1****Chapter Test B****Chapter: Chemical Equations and Reactions**

**Part 1** In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 1. The production of a slightly soluble solid compound in a double-displacement reaction is called the formation of a  
a. gas.  
b. precipitate.  
c. molecular reaction.  
d. halogen.
- \_\_\_\_\_ 2. To balance a chemical equation, it is permissible to adjust the  
a. coefficients.  
b. subscripts.  
c. formula of the products.  
d. number of products.
- \_\_\_\_\_ 3. In a chemical equation, the symbol (aq) indicates that the substance is  
a. melted.  
b. dissolved in water.  
c. an acid.  
d. insoluble.
- \_\_\_\_\_ 4. The tendency for a replacement reaction to occur depends on the  
a. relative tendency of the two elements to lose electrons.  
b. temperature difference.  
c. relative electron loss tendency.  
d. relative tendency of the two elements to lose electrons.
- \_\_\_\_\_ 5. The coefficients in a chemical equation  
a. indicate the number of moles of each substance that react.  
b. show the number of grams of each substance that react.  
c. are the molar masses of the substances.  
d. show the relative electrons that each atom.
- \_\_\_\_\_ 6. If metal X is more reactive than metal Y in the activity series, then  
a. X will replace ions of Y in solution.  
b. Y will replace ions of X in solution.  
c. Y will displace cations only indirectly.  
d. X will react with cold water.

# Holt Study Guide Answer Key

**Terry C. Jones**



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