

Gap Analysis

GOAL: Reduce SecOps' MTTR

Current state	Future state	Gap description	Possible next steps
-MTTR is 5 hours on average -Digging through logs and alerts from different sources -A lot of time spent on false positives -SecOps team is frustrated at how much time wasted	-MTTR is 1 hour -Streamlined process -Less strain on SecOps' individual workloads	-MTTR difference of 80% -Inefficient process -Alerts and logs identify that something maybe happened, but no other context -Too many places to look	-Hire more SecOps people -Aggregate data in one place for analysis -Switch to NetFlow data -Create stricter alerting policies

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web the molar volume will be the volume occupied by one molar mass 32 g of the liquid expressing the density in liters

instead of ml we have $V_m = \frac{32 \text{ g mol}^{-1}}{1.790 \text{ g l}^{-1}} = 0.0405 \text{ l mol}^{-1}$ nonnumber

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web the molar volume of a gas from avogadro's law the volume is directly proportional to the number of moles the important point to keep in mind is that n can be moles of any gas since one mole always contains 6.02×10^{23} molecules and one molecule of any gas at the same temperature will have the same kinetic energy based on this principle it is found

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web apr 22 2020 molar volume of gases 1 counting in moles one mole is defined as 6.02×10^{23} particles this is the avogadro number an unchanging constant whether you are counting hydrogen atoms carbon dioxide molecules or carbonate ions a mole is always 6.02×10^{23} but as different particles have different mass their molar mass varies

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web molar volume of a gas this calculator calculates molar volume of ideal gas at different conditions non standard temperature and pressure

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web mar 12 2018 you can define the concentration of a gas though and that calculation would be exactly the one you've described placing x moles of a gas in a container of volume V yields a concentration $C = \frac{x}{V}$ with units of $\frac{\text{g}}{\text{pu mol over dm}^3}$

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web the molar volume has the SI unit of cubic metres per mole $\text{m}^3 \text{ mol}^{-1}$ although it is more typical to use the units cubic decimetres per mole $\text{dm}^3 \text{ mol}^{-1}$ for gases and cubic centimetres per mole $\text{cm}^3 \text{ mol}^{-1}$ for liquids and solids

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web zero degrees celsius or 273 K is defined as standard temperature and one atmosphere or 760 mm Hg is defined as standard pressure the volume occupied by one mole of a gas is called molar volume therefore the molar volume of a gas at standard temperature and pressure STP is 22.4 l

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web at standard temperature and pressure STP the molar volume V_m is the volume occupied by one mole of a chemical element or a chemical compound it can be calculated by dividing the molar mass M by mass density ρ molar gas volume is one mole of any gas at a specific temperature and pressure has a fixed volume molar volume formula

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web this volume is known as the molar gas volume V_m the volume of a gas may be calculated from its number of moles using volume of gas moles V_m rearranging this moles of gas $\frac{V}{V_m}$

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web jun 8 2023 $V = 1.00 \text{ mol} \times 0.08206 \text{ L atm mol}^{-1} \text{ K}^{-1} \times 273 \text{ K} = 22.4 \text{ L}$ in other words 1 mole of a gas will occupy 22.4 l at stp assuming ideal gas behavior at stp the volume of a gas is only dependent on number of moles of

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web aug 3 2022 volumes of gases and liquids are measured in cubic centimetres cm^3 or cubic decimetres dm^3 or litres l 1 mole of any gas at room temperature and pressure occupies 24 dm^3 $24,000 \text{ cm}^3$ this value is known as the molar volume

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