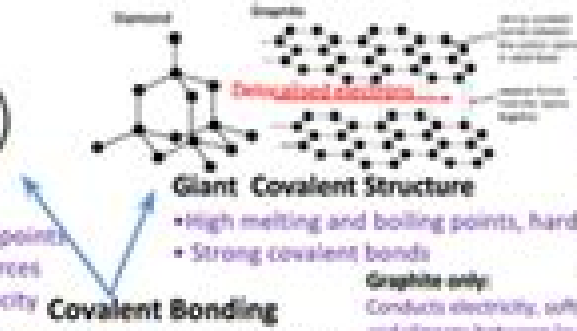
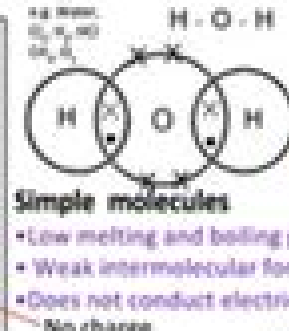




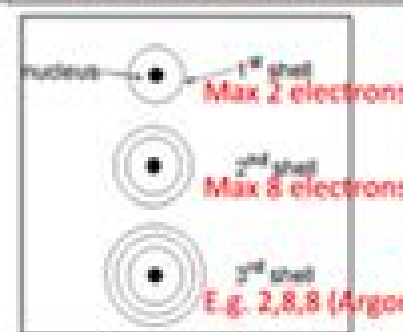
Periodic Table of the Elements

An atom:  
 Number of Protons = Number of Electrons  
 Mass number = Protons + Neutrons

# C2 Chemistry

Periodic table arranges elements by **ATOMIC NUMBER** (proton number/the small one)

The number of outer shell electrons match the group the element is found in.  
 E.g. Lithium 2,1 is a group 1 element.

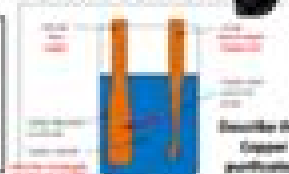


Protons and neutrons both have a mass of 1

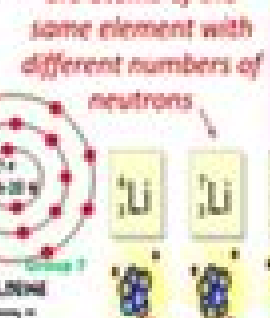
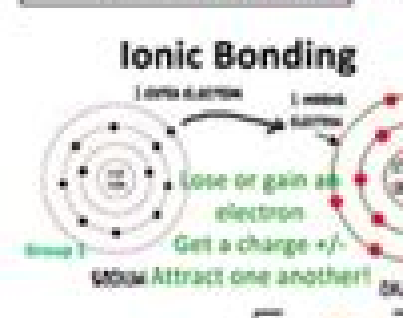
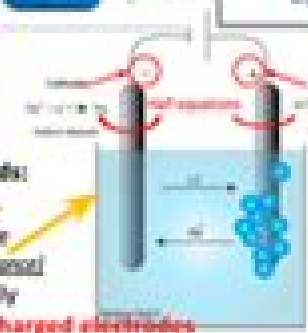
**ISOTOPES** are atoms of the same element with different numbers of neutrons

**OIL RIG**

Oxidation is loss (of electrons)  
 Reduction is gain (of electrons)



**Electrolysis:**  
 Splitting up a substance using electricity



When there are 2 possible products  
 The less reactive forms at the electrode

Electrolysis of Brine

Used for: Disinfectants, Bleach and plastics

Used for: Margarine and HCl

Used for: Bleach, paper and soap

Metal ions only arranged in layers

**ELECTROSTATIC ATTRACTION** between **IONS and ELECTRONS**

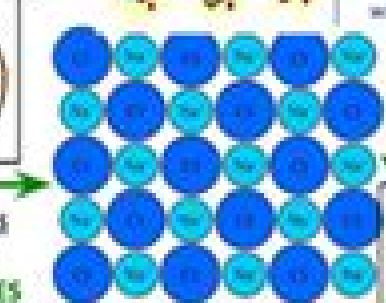
Metal ions slide over one another making them **MALLEABLE**

Metals conduct **HEAT** well

**Giant Lattice Structure**

Always metals with non-metals

+ ION and - ION attracted by **STRONG ELECTROSTATIC FORCES**



As a solid: High melting and Boiling Point

Conducts electricity when melted or dissolved: ions move freely and carry a charge

**NANO-TECHNOLOGY**

Means **REALLY REALLY REALLY** small!!!!

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