

10.1**THE NATURE OF GASES****SECTION REVIEW****Objectives**

- Describe the motion of gas particles according to the kinetic theory
- Interpret gas pressures in terms of kinetic theory

Key Terms

- kinetic energy
- kinetic theory
- gas pressure
- vacuum
- atmospheric pressure
- barometric pressure
- perfect (P₀)
- standard atmospheric (atm)

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are highlighted in this section. Write clearly over the completed work a term, short phrase, or sentence.

- The kinetic theory describes the _____ of particles in matter and the forces of attraction between them. The theory assumes that the volume occupied by a gas is mostly _____, that the particles of gas are _____, move _____ of each other, and are in constant _____ motion. Furthermore, the _____ between particles are elastic so that the total _____ remains constant. The pressure of a gas results from the countless collisions of the gas particles with an object. Barometers are used to measure _____ pressure. Standard conditions are defined as a temperature of _____ and a pressure of _____.
1. _____
 2. _____
 3. _____
 4. _____
 5. _____
 6. _____
 7. _____
 8. _____
 9. _____
 10. _____

Part B True-False

Classify each of these statements as true (always true), AT (sometimes true), ST (sometimes true), or F (false).

- _____ 11. Atmospheric pressure is 760 mm Hg.
- _____ 12. The SI unit of pressure is the pascal.
- _____ 13. Atmospheric pressure decreases as you climb a mountain because the air layer that surrounds the earth thins out.

Nature Of Gases Section Review Key

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