

NAME: _____

CLASS: _____

DATE: _____

1. The molar volume of a gas at STP occupies _____.

- | | | | |
|-------------------------|--------|-------------------------|--------------|
| <input type="radio"/> A | 22.4 L | <input type="radio"/> B | 12 grams |
| <input type="radio"/> C | 0°C | <input type="radio"/> D | 1 kiloPascal |

2. In a double replacement, _____.

- | | | | |
|-------------------------|---|-------------------------|--|
| <input type="radio"/> A | Energy in the form of light or heat is often produced | <input type="radio"/> B | The reactants are generally 2 ionic compounds in aqueous solutions |
| <input type="radio"/> C | The reactants are usually a metal and a nonmetal | <input type="radio"/> D | One of the reactants is often water |

3. If 2.00 L of a gas in a container at room temperature exerts a pressure of 4.00 atm. If the pressure doubles and the temperature remains the same, the volume would become _____.

- | | | | |
|-------------------------|--------|-------------------------|--------|
| <input type="radio"/> A | 16.0 L | <input type="radio"/> B | 1.00 L |
| <input type="radio"/> C | 4.00 L | <input type="radio"/> D | 2.00 L |

4. If 7.00 grams of nitrogen gas are contained in a 5.00 L flask at 462 kPa. What is the temperature of the gas?

- | | | | |
|-------------------------|--------|-------------------------|--------|
| <input type="radio"/> A | 4.12 K | <input type="radio"/> B | 9.07 K |
| <input type="radio"/> C | 11.1 K | <input type="radio"/> D | 1.00 K |

5. Which of the following statements is NOT true about what happens in all chemical reactions?

- | | | | |
|-------------------------|--|-------------------------|--|
| <input type="radio"/> A | The starting substances are called reactants | <input type="radio"/> B | The bonds of reactants are broken and new bonds of products are formed |
| <input type="radio"/> C | The ways in which atoms are joined together change | <input type="radio"/> D | New atoms are formed as products |

National Paper Grade 11 Chemistry Final Exam 2014

Clifford Lane



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