

EOC Practice Test Form A

1. Given $\odot N$, what is $m\widehat{AB}$?

146



2. Which is NOT a step in the construction of the bisector of $\angle CAB$?



- (A) With one point of the compass on A, draw an arc that intersects both rays of the angle at B and C.
- (B) Place the compass on point B and draw an arc.
- (C) Open the compass wider and place it on point C, then draw an arc.
- (D) Draw a ray through point A and the intersection of the arcs drawn from points B and C.

3. If $n \parallel m$, find $m\angle 1 + m\angle 2$.



80

4. The pyramid shown needs to be placed into a box. The volume of the pyramid is 192 mm^3 .



If each box has the same base dimensions as the pyramid, will the pyramid fit into a box with the given height?

Height of box (mm)	Yes	No
10	☐	☒
13	☒	☐
8	☐	☒

Geometry Eoc Practice With Answers

Jicheng Xie

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \quad \Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g}) \quad \Delta h = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g})$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l}) \quad \Delta h = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l}) \quad \Delta h = 1560 \text{ kJ/mole}$
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{ O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta h = 285 \text{ kJ/mole}$

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web thus for reaction 9 2 in which 1 mole of H_2SO_4 and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = 55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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