

50/50

## Experiment 9 Prelaboratory Assignment

### A Volumetric Analysis

Date 03/30

Desk No. \_\_\_\_\_

1. a. What is the titrant for this experiment?

NaOH (base), sodium hydroxide

- b. Is the indicator generally added to the titrant or the analyte in a titration?

The indicator, phenolphthalein, is added to the analyte, the acid.  
It is colorless in an acid but pink in basic solutions.

2. a. What is the primary standard used in this experiment (name and formula)? Define a primary standard.

A primary standard is an accurate mass of a solid substance which is measured, dissolved in water, and then reacted.

↳  $\text{KHC}_8\text{H}_4\text{O}_4$ , potassium hydrogen phthalate

- b. What is the secondary standard used in this experiment (name and formula)? Define a secondary standard.

A secondary standard is the standard remaining solution from after the molar concentration of the titrant is measured.

↳ NaOH, sodium hydroxide

3. Distinguish between a stoichiometric point and an endpoint in an acid-base titration.

reaction is complete  
because amount  
reactants that have  
combined

when the indicator changes  
colors due to the change  
in pH

4. a. When rinsing a buret after cleaning it with soap and water, should the rinse be dispensed through the buret tip or the top opening of the buret? Explain.

The water and rinse should go through the tip. Solution will dispense through and help keep the concentration of the titrant the same.

- b. Experimental Procedure, Part A.4. In preparing the buret for titration, the final rinse is with the NaOH titrant rather than with deionized water. Explain.

The water can alter the values and the NaOH is hygroscopic.

5. Part A.3. Calculate the mass of  $\text{KHC}_8\text{H}_4\text{O}_4$  (molar mass = 204.23 g/mol) that reacts with 15 mL of the 0.15 M NaOH solution prepared in Part A.2. Express this mass  $\text{KHC}_8\text{H}_4\text{O}_4$  to the correct number of significant figures and record the calculation on the Report Sheet.

$$15 \text{ mL} = 0.015 \text{ L}$$

$$0.015 \text{ L (0.15 M)} = 0.00225 \text{ mol NaOH}$$

$$0.00225 \text{ mol NaOH} \cdot \frac{1 \text{ mol KHC}_8\text{H}_4\text{O}_4}{1 \text{ mol NaOH}} = 0.00225 \text{ mol KHC}_8\text{H}_4\text{O}_4$$

$$0.00225 \text{ mol KHC}_8\text{H}_4\text{O}_4 \times \left( \frac{204.23 \text{ g}}{1 \text{ mol}} \right) = \boxed{0.4595 \text{ g}}$$



# General Chemistry Pre Lab Answers

**M Walker**



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## **Table of Contents General Chemistry Pre Lab Answers**

1. Understanding the eBook General Chemistry Pre Lab Answers
  - The Rise of Digital Reading General Chemistry Pre Lab Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry Pre Lab Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an General Chemistry Pre Lab Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from General Chemistry Pre Lab Answers

- Personalized Recommendations
- General Chemistry Pre Lab Answers User Reviews and Ratings
- General Chemistry Pre Lab Answers and Bestseller Lists
- 5. Accessing General Chemistry Pre Lab Answers Free and Paid eBooks
  - General Chemistry Pre Lab Answers Public Domain eBooks
  - General Chemistry Pre Lab Answers eBook Subscription Services
  - General Chemistry Pre Lab Answers Budget-Friendly Options
- 6. Navigating General Chemistry Pre Lab Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - General Chemistry Pre Lab Answers Compatibility with Devices
  - General Chemistry Pre Lab Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of General Chemistry Pre Lab Answers
  - Highlighting and Note-Taking General Chemistry Pre Lab Answers
  - Interactive Elements General Chemistry Pre Lab Answers
- 8. Staying Engaged with General Chemistry Pre Lab Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers General Chemistry Pre Lab Answers
- 9. Balancing eBooks and Physical Books General Chemistry Pre Lab Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection General Chemistry Pre Lab Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine General Chemistry Pre Lab Answers
  - Setting Reading Goals General Chemistry Pre Lab Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of General Chemistry Pre Lab Answers

- Fact-Checking eBook Content of General Chemistry Pre Lab Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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