- According to the Bronsted—Lowry definition, which chemical species can function both as an acid and as a base?
- 17 a CI SO,2

17

17

- c. NH.
- d. HCO<sub>3</sub>
- e. H<sub>3</sub>O\*
- In the reaction CN<sup>-</sup> + H<sub>2</sub>O ←→ HCN + OH<sup>-</sup> which is an acid—base conjugate pair?
  - a. H2O and HCN
  - b. H<sub>2</sub>O and OH
  - c. CN<sup>-</sup> and H<sub>2</sub>O
  - d. HCN and OH
- Given that HX is a stronger Bronsted acid than HY in aqueous solution, which is true of a 1 M solution of NaX
  - a. It is less basic than a 1 M solution of NaY.
  - It is more basic than a 1 M solution of NaY.
  - It yields a neutral solution.
  - It is more concentrated than a 1 M solution of NaY.
- 4. The pH of a 0.03 M HCl solution is
- a. 1.5
  - b. 2.5
  - c. 3.5 d. 12.5
  - 5. The pH of a 1.0 x 10<sup>-3</sup> M Ba(OH)<sub>3</sub> solution at
  - 25°C is a. 2.7
    - a. 2.7 A. 3.0
    - c. 11.0
    - d. 11.3
  - 6. Which is the strongest acid?
- a. HCiO
  - & HClO<sub>2</sub>
  - c. HClO<sub>3</sub>
  - d. HClO<sub>2</sub>
  - 7. Which salt reacts with water (hydrolyzes) to produce a basic solution?
- F & NaC2H3O2
  - A. NaNO
  - c. NH<sub>4</sub>Cl
  - at BaSO.

8. The weakest of the bases listed is

Aeid	Conjugate Base	Ka (Ionization Constant of Acid)	
HCI	CI	100% ionized	
HSO <sub>4</sub>	SO,2	1.2 x 10 <sup>-2</sup>	
H <sub>2</sub> S	HS	5.7x10 <sup>-8</sup>	
HIS	S	1.2×10 <sup>-15</sup>	7

at CE

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- & CN
- c. HS
- d. 52-
- e. SO<sub>2</sub>2-

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Table 1-1	100			-			7	-	



- Which statement is a logical inference from the fact that a 0.10 M solution of potassium acetate KC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>, is less alkaline than a 0.10 M solution of potassium cyanide, KCN?
- A. Hydrocyanic acid is a weaker acid than acetic acid.
  - Hydrocyanic acid is less soluble in water than acetic acid.
  - Cyanides are less soluble than acetates
  - A Acetic soid is a weaker soid than hydrocyanic soid
  - In the titration of 50.0 mL of 0.100 M benzoic acid (a monoprotic acid) with 50.0 mL of 0.100 M NaOH, the properties of the solution at the equivalence point will correspond exactly to the properties of
  - a. a 0.100 M sodium solution.
    - a 0.0500 M sodium hydroxide solution.
    - c. a 0.0500 M benzoic acid solution.
    - d. a 0.0500 M sodium benzoate solution.

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**H Kauffman** 

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