



Molar Volume of Hydrogen

Combining the Gas Laws

Introduction

Airhage have been recuired safety features on new cars since the 1980's and are credited with saring thousands of lives in that time. Airhage contain a compound that decomposes to give reimpen gas upon inspect from a collision. How much nitrogen gas need, be generated to fill an airhage. The amount of gas received to fill any size container can be calculated if we know the notific volume of the gas.

Concepts

Assignative in Law.

· Dahoek bes

· Total gualier.

Midden weikinger.

Background

Associate's law stage that equal volumes of goes contain equal numbers of nodecales under the same conditions of temperature and pressure. It follows, therefore, that all gas samples containing the same number of molecules will occupy the same volume if the temperature and pressure are kept constant. The volume occupied by one node of a gasts called the molecules us are pressured by one node of a gasts called the molecules us sent to this experiment, we will necesse the node column of hydrogen gas at standard temperature and pressure (STP, equal to 273 K and 1 atroit.

The reaction of magnesium nortal with hydrochloric actif Ecuation 13 provides a convenient rorans of generating small reads cumuties of hydrogen in the Iala.

If the reastion is derived out with excess hydrochloric and, the volume of hydrogen gas obtained will depend on the mander of modes of magnesians as well as on the pressure and temperature. The mode volume of hydrogen can be advanted if we recovere the volume excupied by a sample containing a known number of modes of hydrogen. Since the solume will be recovered under blancatury conditions of temperature and pressure, the recovered volume must be corrected to STP conditions before calculating the moder volume.

The relationship among the four gas variables—pressure (P), volume (P), temperature (P) and the number of moles (x)—is expressed in the ideal gas law (Equation 2), where N is a constant called the universal gas constant.

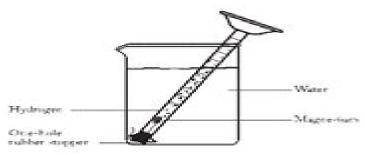
Elyanothera 2.

The total gas law reduces to Ex aution 3, the condined gas law, if the number of number of gas is constant. The conditioning gas law can be used to calculate the volume (t_{ij}) of a gas at STP (T_{ij}) and P_{ij}) from the order of the number of laboratory conditions (T_{ij}) and P_{ij}). In using either the ideal gas law or the condition igns law, remember that temperature must be always be expressed to arise of kelvins (A,b) on the shocket temperature react.

$$\frac{P_z F_z}{\gamma_z} = \frac{P_z F_z}{\gamma_z}$$
 Equation 3

Highragen gas will be collected by the displacement of materin an inversel graduated cylinder using the apparatus above in Figure 1. The total pressure of the gas in the cylinder will be edual to the harmouric tair) pressure. However, the gas in the cylinder will may be pure hydrogen. The gas will also contain water vapor due to the evaporation of the water molecules over which it is being collected. According to Dalum's law, the total pressure of the gas will be equal to the partial pressure of hydrogen plus the partial pressure of outer vapor (Rouation 4). The vapor pressure of water includes pressure for Table 1).

$$P_{+,+} = P_{+,-} + P_{+,+}$$
 Equation 4.



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