

EXERCISES FOR CHAPTER I

Part I. Exercises with solution.

1. How can we classify different kinds of crystal lattice? What is the relation between structure and physical properties?
 - Classifying types of crystal lattice depends on:
 - + Distance between the lattice nodes
 - + Distance between the lattice nodes comparing to the distance between atoms in a unit cell.
 - How do these factor affect to the physical properties:
 - + Island structure (CO_2 , XeF_4 , Ar, H_2O , $\text{K}_2[\text{TiCl}_6]$): the destruction of chemical bonding between the lattice nodes is easier than the destruction of bonding between atoms in a lattice nodes.
 - + Chain structure: covalent bonding in one dimension, octahedron, tetrahedron.
 - Van de Waals fore: the lattice structure is easy to be stripped in to yarn.
 - pi bonding: good conductor, following the direction of chain.
 - features: not crystal-clear, following the direction of chain.
 - Layer crystal structure: covalent bonding in 2 dimensions.
 - Van de Waals force: easy so strip, not hard.
 - Pi bonding: good conductor.
 - Not crystal clear, low hardness, low proportion.
 - Coordination structure: lattice nodes are atoms, singular ions liking to each other by strong forces: metallic bonding, covalent bonding, ionic bonding.
2. What is the critical temperature, the critical volume? Find the phase diagram of CO_2 . Based on the phase diagram of the movement of CO_2 explain the transformation of CO_2 in this schematic. Point three (triple point) on meaningful schema like? Mentioned application of supercritical CO_2 .
 At normal pressure, chemical gas at a specified temperature. Temperature temperature is called liquefaction. In contrast, the liquid at that temperature and evaporation, so that temperature is the boiling point of the liquid.
 However, raising the temperature of liquefaction (or boiling) because the pressure has a certain limit, through which the liquid temperature can not exist despite public pressure.
 The maximum temperature is called the critical temperature (TTH) and pressure needed to liquefy the gas at that temperature called the critical pressure (PTH). Molar volume of a gas at the critical temperature and critical pressure is called the critical volume. In extreme conditions, the volume of gases and liquids equal to where gas and liquids with the same mass ratio.
3. All metals have high melting temperature and boiling temperature. Is this statement right or wrong? Explain?

| Metals | Valence configuration | Melting temperature($^{\circ}\text{C}$) | Boiling temperature($^{\circ}\text{C}$) |
|--------|-----------------------|---|---|
| Hg | $6s^2$ | -38,89 | 356,66 |
| W | $5d^46s^2$ | 3420 | 5680 |
| Pb | $6s^26p^2$ | 327,4 | 1745 |
| K | $4s^1$ | 63,55 | 761 |
| Tl | $6s^26p^1$ | 304 | 1475 |

Multiple Choice Questions On Inorganic Chemistry

Marco Cascella



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