

## Chapter 19 Electrochemistry

### Practice questions. Answers

1. What element is being reduced in the following redox reaction?



2. What element is being oxidized in the following redox reaction?



3. What element is being oxidized in the following redox reaction?



4. What element is being reduced in the following redox reaction?



5. Balance the following redox reaction if it occurs in acidic solution. What are the coefficients in front of  $\text{H}_2\text{C}_2\text{O}_4$  and  $\text{H}_2\text{O}$  in the balanced reaction?



$\text{H}_2\text{C}_2\text{O}_4 = 5, \text{H}_2\text{O} = 8$

6. Balance the following redox reaction if it occurs in acidic solution. What are the coefficients in front of  $\text{H}^+$  and  $\text{Fe}^{3+}$  in the balanced reaction?



$\text{H}^+ = 8, \text{Fe}^{3+} = 5$

7. Determine the cell notation for the redox reaction given below.



$\text{Fe}(\text{s}) \mid \text{Fe}^{3+}(\text{aq}) \parallel \text{Cl}_2(\text{g}) \mid \text{Cl}^-(\text{aq}) \mid \text{Pt}$

8. Use the standard half-cell potentials listed below to calculate the standard cell potential for the following reaction occurring in an electrochemical cell at  $25^\circ\text{C}$ . (The equation is balanced.)



# Fall 2012 Electrochemistry Test Answers

**Richard H. Langley, John Moore**



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