

STD XII - CHEMISTRY

Time : 3 hrs.]

MARCH - 2013

[Marks : 70

SECTION - I

Q. 1. Select and write the most appropriate answers from the given alternatives for each sub-question :

[7]

- (i) In body centred cubic structure the space occupied is about _____.
 (a) 68% (b) 53% (c) 38% (d) 32%
- (ii) For a gaseous reaction the unit of rate of reaction is _____.
 (a) L atm s^{-1} (b) $\text{atm mol}^{-1} \text{s}^{-1}$ (c) atm s^{-1} (d) mol s
- (iii) Which of the following compounds contains S = O as well as S = S bonds?
 (a) Sulphuric acid (b) Thiosulphuric acid
 (c) Sulphurous acid (d) Thiosulphurous acid
- (iv) Which of the following solutions shows maximum depression in freezing point?
 (a) 0.5 M Li_2SO_4 (b) 1 M NaCl (c) 0.5 M $\text{Al}_2(\text{SO}_4)_3$ (d) 0.5 M BaCl_2
- (v) For a chemical reaction $dS = -0.035 \text{ kJ/K}$ and $dH = -20 \text{ kJ}$. At what temperature does the reaction turn nonspontaneous?
 (a) 5.14 K (b) 57.14 K (c) 571.4 K (d) 5714.0 K
- (vi) The standard e.m.f. of the following cell is 0.463 V

$$\text{Cu}_{(s)} \mid \text{Cu}_{(aq)}^{++} (1\text{M}) \parallel \text{Ag}_{(aq)}^{++} (1\text{M}) \mid \text{Ag}_{(s)}$$
 . What is the standard potential of Cu electrode?
 (a) 1.137 V (b) 0.337 V (c) 0.463 V (d) -0.463 V
- (vii) Fe_2O_3 is reduced to spongy iron near the top of blast furnace by _____.
 (a) H_2 (b) CaO (c) SiO_2 (d) CO

Q. 2. Answer any SIX of the following :

[12]

- (i) Distinguish between crystalline solid and amorphous solid.
- (ii) State Kohlrausch Law and write mathematical expression of molar conductivity of the given solution at infinite dilution.
- (iii) Write cell reactions in lead storage battery during discharge.
- (iv) Draw structures and write geometry of PCl_3 and PCl_5 .
- (v) Prove that $\Delta H = \Delta U + \Delta nRT$. What is the condition under which $\Delta U = \Delta H$?
- (vi) Mention names and formulae of two ores of aluminium.
- (vii) Derive the relationship between relative lowering of vapour pressure and molar mass of nonvolatile solute.
- (viii) What is pseudo first order reaction? Give one example of it.

Q. 3. Answer any THREE of the following :

[9]

- (i) Calculate the mole fraction and molality of HNO_3 in a solution containing 12.2% HNO_3 . (Given - atomic masses : H = 1, N = 14, O = 16)
- (ii) Consider the reaction -



What are the values of the following at the same time ?

- (a) $-\frac{d[\text{I}^-]}{dt}$ (b) $-\frac{d[\text{S}_2\text{O}_8^{2-}]}{dt}$ (c) $-\frac{d[\text{I}_3^-]}{dt}$

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, 1979-03 *HSC Standard English* Barry Spurr, Lloyd Cameron, 2009 This guide contains comprehensive summary and discussion of all 44 prescribed texts in the HSC Standard English course plus a list of key issues to consider in each chapter related to the relevant syllabus area, helpful advice on how to read different types of texts, plot outlines, character discussion and interpretations. [Handbook of In Vivo Chemistry in Mice](#) Katsunori Tanaka, Kenward Vong, 2020-04-27 Provides timely

comprehensive coverage of in vivo chemical reactions within live animals. This handbook summarizes the interdisciplinary expertise of both chemists and biologists performing in vivo chemical reactions within live animals. By comparing and contrasting currently available chemical and biological techniques, it serves not just as a collection of the pioneering work done in animal-based studies but also as a technical guide to help readers decide which tools are suitable and best for their experimental needs. The Handbook of In Vivo Chemistry in Mice: From Lab to Living System introduces readers to general information about live animal experiments and detection methods commonly used for these animal models. It focuses on chemistry-based techniques to develop selective in vivo targeting methodologies as well as strategies for in vivo chemistry and drug release. Topics include currently available mouse models, biocompatible fluorophores, radionuclides for

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