

# Lab 11

## Gravimetric Analysis

### Pre-Lab Questions

1. Write a balanced equation for the precipitation of calcium carbonate from potassium carbonate and calcium chloride.  
$$\text{K}_2\text{CO}_3 + \text{CaCl}_2 \rightarrow \text{CaCO}_3 + 2\text{KCl}$$
2. Using this balanced equation, determine the limiting reactant if 15 grams of calcium chloride was reacted with 15 grams of potassium carbonate.

Using your answer for Question 2, determine the mass of potassium carbonate needed to fully precipitate all the calcium from a 25 mL sample of 15% calcium chloride.

# Gravimetric Analysis Prelab Answers

**László Erdey, Ilona Buzás, Gyula Svehla**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located behind the author names.

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**Academic Computing**, 1987     *Investigation in General Chemistry* Alice S. Cohen, 1978     **Book of Abstracts** American Chemical Society. Meeting, American Chemical Society, 1974     **Laboratory Experiments for General Chemistry** Harold R. Hunt, Toby F. Block, 1994     General Chemistry Hunt, 1976-06-07     **Gravimetric Analysis** László Erdey, 1963     *Gravimetric Analysis* Laszlo Erdey, 1963     Gravimetric analysis László Erdey, Ilona Buzás, Gyula Svehla, 1963

Gravimetric Analysis László Erdey, 2013-10-22 Gravimetric Analysis Part III describes the experimental procedures for the gravimetric analysis of various compounds This book is composed of 13 chapters that also present sample preparation protocols The first four chapters survey the steps for halogen compound determination The succeeding chapters provide the procedures for gravimetric determination of cyanide thiocyanate ions sulfur nitrogen phosphorus carbon silicon and boron The final chapter considers other aspects of gravimetric experiments including apparatus cleaning reagents and numerical calculation of the result This book will prove useful to analytical and inorganic chemists teachers and students in the allied fields     *Introductory Titrimetric and Gravimetric Analysis* Evelyn M. Rattenbury, 2016-06-06 Introductory Titrimetric and Gravimetric Analysis discusses the different types of titration and the weighing of different solutions in solid form Coverage is made on acid base titration argentometric titrations and oxidation reduction titrations Iodometric titrations and complexometric titrations are also explained Extensive discussion on each of the titration method along with some examples and laboratory experiments is given The process of weight measurement of damp powder is one example of the experiments The book is a manual that guides a student to the correct ways of conducting an experiment made on such solutions as sodium hydroxide using hydrochloric acid and oxalic acid Outcome of such experiments in terms of composition weight of solutions and measurement of pressure in certain environment is tabulated and briefly explained Logarithms and antilogarithms are included at the end of the book The text will serve as a good laboratory manual for students preparing for science examination as well as for chemists and chemical engineers     Gravimetric Analysis László Erdey, 2013-09-03 Analytical Chemistry Volume 7 Gravimetric Analysis Part II describes the experimental procedures for the gravimetric analysis of Groups I to V cations This book is composed of 43 chapters that also present sample preparation separation and precipitation protocols The first six chapters include Group I cations such as silver lead mercury copper bismuth and cadmium followed by chapters on Group II cations including arsenic antimony tin germanium gold platinum selenium and

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