

## Prescribing Prescriptions for Chemical Reactions

### Directions:

1. Write the balanced chemical equation and complete the following questions.
2. For the reaction type = Answer from scratch effect, Decomposition (DC), Single Replacement (SR), Double Replacement (DR) and Complex Combustion (CC). Also add the additional classification of Endothermic Reactions (ER) or Exothermic (EX) if appropriate.

- a.  $\text{NaOH}_{(aq)} + \text{HCl}_{(aq)} \rightarrow \text{NaCl}_{(aq)} + \text{H}_2\text{O}_{(l)}$  Reaction Type(s): Double Replacement
- a.  $\text{CaCO}_3_{(s)} \rightarrow \text{CaO}_{(s)} + \text{CO}_2_{(g)}$  Reaction Type(s): Decomposition
- a.  $\text{Zn}_{(s)} + \text{HCl}_{(aq)} \rightarrow \text{ZnCl}_2_{(aq)} + \text{H}_2_{(g)}$  Reaction Type(s): Single Replacement
- a.  $\text{CH}_4_{(g)} + 2\text{O}_2_{(g)} \rightarrow \text{CO}_2_{(g)} + 2\text{H}_2\text{O}_{(l)}$  Reaction Type(s): Combustion
- a.  $\text{Fe}_{(s)} + \text{CuSO}_4_{(aq)} \rightarrow \text{FeSO}_4_{(aq)} + \text{Cu}_{(s)}$  Reaction Type(s): Single Replacement
- a.  $\text{H}_2\text{SO}_4_{(aq)} + \text{NaOH}_{(aq)} \rightarrow \text{Na}_2\text{SO}_4_{(aq)} + \text{H}_2\text{O}_{(l)}$  Reaction Type(s): Double Replacement
- a.  $\text{KClO}_3_{(s)} \rightarrow \text{KCl}_{(s)} + \text{O}_2_{(g)}$  Reaction Type(s): Decomposition
- a.  $\text{HCl}_{(aq)} + \text{NaOH}_{(aq)} \rightarrow \text{NaCl}_{(aq)} + \text{H}_2\text{O}_{(l)}$  Reaction Type(s): Double Replacement
- a.  $\text{Al}_{(s)} + \text{CuSO}_4_{(aq)} \rightarrow \text{Al}_2(\text{SO}_4)_3_{(aq)} + \text{Cu}_{(s)}$  Reaction Type(s): Single Replacement
- a.  $\text{H}_2\text{O}_{(l)} \rightarrow \text{H}_2_{(g)} + \text{O}_2_{(g)}$  Reaction Type(s): Decomposition
- a.  $\text{CH}_4_{(g)} + 2\text{O}_2_{(g)} \rightarrow \text{CO}_2_{(g)} + 2\text{H}_2\text{O}_{(l)}$  Reaction Type(s): Combustion
- a.  $\text{Fe}_{(s)} + \text{CuSO}_4_{(aq)} \rightarrow \text{FeSO}_4_{(aq)} + \text{Cu}_{(s)}$  Reaction Type(s): Single Replacement
- a.  $\text{H}_2\text{SO}_4_{(aq)} + \text{NaOH}_{(aq)} \rightarrow \text{Na}_2\text{SO}_4_{(aq)} + \text{H}_2\text{O}_{(l)}$  Reaction Type(s): Double Replacement
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- a.  $\text{Al}_{(s)} + \text{CuSO}_4_{(aq)} \rightarrow \text{Al}_2(\text{SO}_4)_3_{(aq)} + \text{Cu}_{(s)}$  Reaction Type(s): Single Replacement

# Lab 25 Equation Writing And Predicting Products

## Answer Key

**Lingsheng Yao**



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