

# Acids and Bases

MERCK

Acid				Base			
Strength of Acid	Ka	Name	Formula	Formula	Name	Strength of Base	
Strongest	Large	Perchloric acid	HClO <sub>4</sub>	ClO <sub>4</sub> <sup>-</sup>	Perchlorate ion	Weakest	
	3.2×10 <sup>8</sup>	Hydroiodic acid		I <sup>-</sup>	Iodide		
	1.0×10 <sup>9</sup>	Hydrobromic acid	HBr	Br <sup>-</sup>	Bromide		
	1.3×10 <sup>4</sup>	Hydrochloric acid	HCl	Cl <sup>-</sup>	Chloride		
	1.0×10 <sup>3</sup>	Sulfuric acid	H <sub>2</sub> SO <sub>4</sub>	HSO <sub>4</sub> <sup>-</sup>	Hydrogen sulfate ion		
	2.4×10 <sup>1</sup>	Nitric acid	HNO <sub>3</sub>	NO <sub>3</sub> <sup>-</sup>	Nitrate ion		
	.....	Hydronium ion	H <sub>3</sub> O <sup>+</sup>	H <sub>2</sub> O	Water		
	5.4×10 <sup>-2</sup>	Oxalic acid	HO <sub>2</sub> C <sub>2</sub> O <sub>2</sub> H	HO <sub>2</sub> C <sub>2</sub> O <sub>2</sub> <sup>-</sup>	Hydrogen oxalate ion		
	1.3×10 <sup>-2</sup>	Sulfurous acid	H <sub>2</sub> SO <sub>3</sub>	HSO <sub>3</sub> <sup>-</sup>	Hydrogen sulfite ion		
	1.0×10 <sup>-2</sup>	Hydrogen sulfate ion	HSO <sub>4</sub> <sup>-</sup>	SO <sub>4</sub> <sup>2-</sup>	Sulfate ion		
	7.1×10 <sup>-3</sup>	Phosphoric acid	H <sub>3</sub> PO <sub>4</sub>	H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>	Dihydrogen phosphate ion		
	7.2×10 <sup>-4</sup>	Nitrous acid	HNO <sub>2</sub>	NO <sub>2</sub> <sup>-</sup>	Nitrite ion		
	6.6×10 <sup>-4</sup>	Hydrofluoric acid	HF	F <sup>-</sup>	Fluoride ion		
	1.8×10 <sup>-4</sup>	Methanoic acid	HCO <sub>2</sub> H	HCO <sub>2</sub> <sup>-</sup>	Methanoate ion		
	6.3×10 <sup>-5</sup>	Benzoic acid	C <sub>6</sub> H <sub>5</sub> COOH	C <sub>6</sub> H <sub>5</sub> COO <sup>-</sup>	Benzoate ion		
	5.4×10 <sup>-5</sup>	Hydrogen oxalate ion	HO <sub>2</sub> C <sub>2</sub> O <sub>2</sub> <sup>-</sup>	O <sub>2</sub> C <sub>2</sub> O <sub>2</sub> <sup>2-</sup>	Oxalate ion		
	1.8×10 <sup>-5</sup>	Ethanoic acid	CH <sub>3</sub> COOH	CH <sub>3</sub> COO <sup>-</sup>	Ethanoate (acetate) ion		
	4.4×10 <sup>-7</sup>	Carbonic acid	H <sub>2</sub> CO <sub>3</sub>	HCO <sub>3</sub> <sup>-</sup>	Hydrogen carbonate ion		
	1.1×10 <sup>-7</sup>	Hydrosulfuric acid	H <sub>2</sub> S	HS <sup>-</sup>	Hydrogen sulfide ion		
	6.3×10 <sup>-8</sup>	Dihydrogen phosphate ion	H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>	HPO <sub>4</sub> <sup>2-</sup>	Hydrogen phosphate ion		
	6.2×10 <sup>-8</sup>	Hydrogen sulfite ion	HS <sup>-</sup>	SO <sub>3</sub> <sup>2-</sup>	Sulfite ion		
	2.9×10 <sup>-8</sup>	Hypochlorous acid	HClO	ClO <sup>-</sup>	Hypochlorite ion		
	6.2×10 <sup>-10</sup>	Hydrocyanic acid	HCN	CN <sup>-</sup>	Cyanide ion		
	5.8×10 <sup>-10</sup>	Ammonium ion	NH <sub>4</sub> <sup>+</sup>	NH <sub>3</sub>	Ammonia		
	5.8×10 <sup>-10</sup>	Boric acid	H <sub>3</sub> BO <sub>3</sub>	H <sub>2</sub> BO <sub>3</sub> <sup>-</sup>	Dihydrogen borate		
	4.7×10 <sup>-11</sup>	Hydrogen carbonate ion	HCO <sub>3</sub> <sup>-</sup>	CO <sub>3</sub> <sup>2-</sup>	Carbonate ion		
	4.2×10 <sup>-13</sup>	Hydrogen phosphate ion	HPO <sub>4</sub> <sup>2-</sup>	PO <sub>4</sub> <sup>3-</sup>	Phosphate ion		
	1.8×10 <sup>-13</sup>	Dihydrogen borate ion	H <sub>2</sub> BO <sub>3</sub> <sup>-</sup>	HBO <sub>3</sub> <sup>2-</sup>	Hydrogen borate ion		
	1.3×10 <sup>-13</sup>	Hydrogen sulfide ion	HS <sup>-</sup>	S <sup>2-</sup>	Sulfide ion		
	1.6×10 <sup>-14</sup>	Hydrogen borate ion	HBO <sub>3</sub> <sup>2-</sup>	BO <sub>3</sub> <sup>3-</sup>	Borate ion		
Weakest	.....	water	H <sub>2</sub> O	OH <sup>-</sup>	Hydroxide	Strongest	



# Lab 16 1 Creating Acid Base Strength Table

**ML Morrison**



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popular and counter culture to denote a biker appearance or attitude it is most notably featured in the 1969 film easy rider  
sometimes born to be wild is described as the first heavy metal song and the second verse lyric heavy metal thunder marks  
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