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Use the References to access important values if needed for this question.

For the reaction



$$\Delta H^\circ = -114.4 \text{ kJ and } \Delta S^\circ = -128.9 \text{ J/K}$$

The maximum amount of work that could be done when 1.54 moles of HCl(g) react at 279 K, 1 atm is kJ.

Assume that ΔH° and ΔS° are independent of temperature.

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A) For
 $\text{CH}_4\text{(g)}$
 ΔH°
The eq
Assume

B) For
 $\text{N}_2\text{(g)}$ +
 ΔH°
The eq
Assume

C) For t
 2HBr(g)
 ΔH°
The equ
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Ilrn Work Answers

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