

Hewlett-Packard Professional Certification Assessment

Examination 1017 has six to sixteen following problems. Use all appropriate and proper scientific units. Use scientific notation for answers requiring scientific notation in the answer blank.

1. Find the standard enthalpy change for the reaction $\text{C}_{10}\text{H}_8 + 12\text{O}_2 \rightarrow 10\text{CO}_2 + 8\text{H}_2\text{O}$
 of carbon (s) $\Delta H_{\text{f}}^\circ = 0 \text{ kJ/mol}$ $\Delta H_{\text{f}}^\circ = 0 \text{ kJ/mol}$ $\Delta H_{\text{f}}^\circ = 0 \text{ kJ/mol}$
 and C_{10}H_8 $\Delta H_{\text{f}}^\circ = 101.32 \text{ kJ/mol}$ $\Delta H_{\text{f}}^\circ = 101.32 \text{ kJ/mol}$

a. $\Delta H_{\text{rxn}}^\circ = 101.32 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 101.32 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 101.32 \text{ kJ/mol}$
 b. $\Delta H_{\text{rxn}}^\circ = 101.32 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 101.32 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 101.32 \text{ kJ/mol}$

c. $\Delta H_{\text{rxn}}^\circ = 101.32 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 101.32 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 101.32 \text{ kJ/mol}$

2. The standard enthalpy change for the formation of aluminum oxide and aluminum hydroxide are $\Delta H_{\text{f}}^\circ = -1675.7 \text{ kJ/mol}$ $\Delta H_{\text{f}}^\circ = -1675.7 \text{ kJ/mol}$ $\Delta H_{\text{f}}^\circ = -1675.7 \text{ kJ/mol}$
 $\Delta H_{\text{f}}^\circ = -1675.7 \text{ kJ/mol}$ $\Delta H_{\text{f}}^\circ = -1675.7 \text{ kJ/mol}$ $\Delta H_{\text{f}}^\circ = -1675.7 \text{ kJ/mol}$
 Calculate the standard enthalpy change for the reaction $2\text{Al} + 3\text{H}_2\text{O} \rightarrow \text{Al}_2\text{O}_3 + 3\text{H}_2$

a. $\Delta H_{\text{rxn}}^\circ = 1675.7 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 1675.7 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 1675.7 \text{ kJ/mol}$
 b. $\Delta H_{\text{rxn}}^\circ = 1675.7 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 1675.7 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 1675.7 \text{ kJ/mol}$

c. $\Delta H_{\text{rxn}}^\circ = 1675.7 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 1675.7 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 1675.7 \text{ kJ/mol}$
 d. $\Delta H_{\text{rxn}}^\circ = 1675.7 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 1675.7 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 1675.7 \text{ kJ/mol}$

e. $\Delta H_{\text{rxn}}^\circ = 1675.7 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 1675.7 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 1675.7 \text{ kJ/mol}$

3. The Gibbs free energy of formation of carbon monoxide, carbon dioxide, and hydrogen, listed from top to bottom are $\Delta G_{\text{f}}^\circ = 137.16 \text{ kJ/mol}$ $\Delta G_{\text{f}}^\circ = 137.16 \text{ kJ/mol}$ $\Delta G_{\text{f}}^\circ = 137.16 \text{ kJ/mol}$

Calculate the standard enthalpy change for the reaction $2\text{CO} + \text{O}_2 \rightarrow 2\text{CO}_2$

a. $\Delta H_{\text{rxn}}^\circ = 137.16 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 137.16 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 137.16 \text{ kJ/mol}$
 b. $\Delta H_{\text{rxn}}^\circ = 137.16 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 137.16 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 137.16 \text{ kJ/mol}$

c. $\Delta H_{\text{rxn}}^\circ = 137.16 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 137.16 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 137.16 \text{ kJ/mol}$
 d. $\Delta H_{\text{rxn}}^\circ = 137.16 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 137.16 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 137.16 \text{ kJ/mol}$

e. $\Delta H_{\text{rxn}}^\circ = 137.16 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 137.16 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 137.16 \text{ kJ/mol}$
 f. $\Delta H_{\text{rxn}}^\circ = 137.16 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 137.16 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 137.16 \text{ kJ/mol}$

g. $\Delta H_{\text{rxn}}^\circ = 137.16 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 137.16 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 137.16 \text{ kJ/mol}$

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Michael Brown



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