

will dissolve to the same extent? Is the solubility of the salt changed?

Model 1 – Dissolving a Salt in 500 mL of Solution

Note: Proceed to Question 1. Do not fill in these tables until instructed to do so.



Beaker 1

	Ca^{2+}	OH^{-}
Initial	0 M	0 M
Change		
Equilibrium		

Beaker 2

	Ca^{2+}	OH^{-}
Initial	0 M	0.200 M
Change		
Equilibrium		

Beaker 3

	Ca^{2+}	OH^{-}
Initial	0.200 M	0 M
Change		
Equilibrium		

Beaker 4

	Ca^{2+}	OH^{-}
Initial	0 M	0.500 M
Change		
Equilibrium		

1. The solubility of what salt is being investigated in the four beakers in Model 1?



Extension Questions Pogil Ions

CL Gary



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