

**Purpose:** To study the ideal gas law and apply the first law of thermodynamics to analyze behavior of gas. Go to <http://phet.colorado.edu/en/simulation/gas-properties> and run the simulation

**Procedure****PART A: THE NATURE OF THE IDEAL GAS LAW**

1: What happens to the pressure in the chamber if you heat the gas?

2: What happens to the temperature in the chamber if you compress the gas?

3: Locate the "Constant Parameter" section of the on-screen control panel. Lock the temperature. Use "Scubie" to slowly compress the gas. What happens to temperature, and what action is taken (by the simulation) to maintain constant temperature?

4: Lock the pressure. (Doing so releases the lock on temperature.) Add heat. What happens to pressure, and what action is taken (by the simulation) to maintain constant pressure?

**PART B: Kinetic Molecular Theory and the First Law of Thermodynamics**

For each of the following four steps record the average pressure (in atm) and the volume

**Table of Pressure and Volume for each step**

Step	1	2	3	4
Pressure (atm)				
Volume ( $\text{m}^3$ )				

1. Add 300 molecules of the heavy species to the chamber and let the temperature and pressure come to equilibrium.
  - a. Using Avogadro's number express the amount of gas in terms of moles.
  - b. Calculate the internal energy ( $U$ ) of the heavy oxygen gas in joules .
  - c. The container is rectangular in shape with an area of  $2.81 \times 10^{-17} \text{ m}^2$ . Using the ruler tool (in measurement tools) measure the interior length of the container and report measurement in meters.  
Length: \_\_\_\_\_ m
  - d. Calculate the volume of the container (in  $\text{m}^3$ )
  - e. Create different volumes and record the average pressure (in atm) and the volume in the table above for step 1.

# Exploring Gas Laws Lab Answers Phet

**Michael Hoelscher**



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