

Worksheet #3—Stoichiometry

Directions: Solve the following stoichiometry problems. Assume all gases are at STP. Don't forget to write out your balanced chemical equation. Make sure to show ALL work to receive FULL CREDIT!

1. Calculate the number of moles of oxygen required to react completely with 14.8 mol Al in the following equation: $4\text{Al}_{(s)} + 3\text{O}_{2(g)} \rightarrow 2\text{Al}_2\text{O}_{3(s)}$

$$\frac{14.8 \text{ moles Al}}{4 \text{ mole Al}} \times \frac{3 \text{ moles of O}_2}{1 \text{ mole Al}} = \underline{11.1 \text{ mol O}_2}$$

3. If 20.0g grams of zinc react with excess hydrochloric acid, how many grams of zinc chloride (ZnCl_2) are produced?
 $\text{Zn}_{(s)} + 2\text{HCl}_{(aq)} \rightarrow \text{ZnCl}_{2(aq)} + \text{H}_{2(g)}$

$$\frac{20 \text{ g Zn}}{65 \text{ g Zn}} \times \frac{1 \text{ mol Zn}}{1 \text{ mol Zn}} \times \frac{1 \text{ mole ZnCl}_2}{1 \text{ mole Zn}} \times \frac{136 \text{ g ZnCl}_2}{1 \text{ mole ZnCl}_2} = \underline{41.8 \text{ g ZnCl}_2}$$

5. How many grams of copper are required to replace 4.00 grams of silver nitrate which are dissolved in water?
 $\text{Cu}_{(s)} + 2\text{AgNO}_{3(aq)} \rightarrow \text{Cu(NO}_3)_2(aq) + 2\text{Ag}_{(s)}$

$$\frac{4.0 \text{ g AgNO}_3}{170 \text{ g AgNO}_3} \times \frac{1 \text{ mol AgNO}_3}{170 \text{ g AgNO}_3} \times \frac{2 \text{ mol AgNO}_3}{1 \text{ mol Cu}} \times \frac{63.5 \text{ g Cu}}{1 \text{ mole Cu}} = \underline{3.0 \text{ g Cu}}$$

2. Calculate the moles of oxygen produced from the decomposition of 0.61 moles of potassium chlorate (KClO_3).



$$\frac{0.61 \text{ moles KClO}_3}{2 \text{ mole KClO}_3} \times \frac{3 \text{ moles of O}_2}{1 \text{ mole KClO}_3} = \underline{0.915 \text{ mol O}_2}$$

4. How many grams of chlorine gas must be reacted with excess sodium iodide (NaI) if 10.0 grams of sodium chloride (NaCl) are needed?
 $2\text{NaI}_{(aq)} + \text{Cl}_{2(g)} \rightarrow 2\text{NaCl}_{(aq)} + \text{I}_{2(g)}$

$$\frac{10 \text{ g NaCl}}{58.5 \text{ g NaCl}} \times \frac{1 \text{ mol NaCl}}{58.5 \text{ g NaCl}} \times \frac{2 \text{ mol NaCl}}{1 \text{ mole Cl}_2} \times \frac{71 \text{ g Cl}_2}{1 \text{ mole Cl}_2} = \underline{6.07 \text{ g Cl}_2}$$

6. How many moles of hydrogen gas are needed to react with 15.1 g chlorine gas to produce hydrogen chloride gas?
 $\text{H}_{2(g)} + \text{Cl}_{2(g)} \rightarrow 2\text{HCl}_{(g)}$

$$\frac{15.1 \text{ g Cl}_2}{71 \text{ g Cl}_2} \times \frac{1 \text{ mole Cl}_2}{71 \text{ g Cl}_2} \times \frac{1 \text{ mole H}_2}{1 \text{ mole Cl}_2} = \underline{0.21 \text{ moles H}_2}$$

Holt Chemistry Problem Solving Work Answers Stoichiometry

James E. Brady, Larry Peck



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Chemistry James E. Brady,John R. Holum,1995-12-29 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping readers develop skills in analytical thinking and problems solving **General Chemistry - Principles and Structure** James E. Brady,Larry Peck,1986 **Modern University Chemistry, Custom Pub** Norbert T. Porile,1993 *Journal of Chemical Education* ,1926 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers *Chemistry* John W. Moore,Conrad L. Stanitski,Peter C. Jurs,2002 CHEMISTRY THE MOLECULAR SCIENCE is intended to help students develop a broad overview of chemistry and chemical reactions an understanding of the most important concepts and models that chemists and those in chemistry related fields use an appreciation of the many ways chemistry impacts our daily lives the ability to apply the facts concepts and models of chemistry appropriately to new situations in chemistry other sciences and engineering and to other disciplines **Chemistry, Student Study Guide** James E. Brady,Fred Senese,2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH₄ from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement Forthcoming Books Rose Arny,2003-04 Modern Chemistry Nicholas D. Tzimopoulos,1993 *Directory of Distance Learning Opportunities* Modoc Press, Inc.,2003-02-28 This book provides an overview of current K 12 courses and programs offered in the United States as correspondence study or via such electronic delivery systems as satellite cable or the Internet The Directory includes over 6 000 courses offered by 154 institutions or distance learning consortium members Following an introduction that describes existing practices and delivery methods the Directory offers three indexes Subject Index of Courses Offered by Level Course

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