





Stoichiometry Question

Consider the synthesis of water produced in the Model. The reaction container is filled with 10.0 g of H_2 and 1.0 g of O_2 .

- a. Which reactant is the degree of progress in the limiting reactant in this case? Show your work. Also indicate that you are given standard quantities in units such as grams, not moles.

$$10.0 \text{ g } H_2 \left(\frac{1 \text{ mol } H_2}{2.016 \text{ g}} \right) = 4.96 \text{ mol } H_2$$

$$1.0 \text{ g } O_2 \left(\frac{1 \text{ mol } O_2}{32.00 \text{ g}} \right) = 0.031 \text{ mol } O_2$$

	$2 H_2 +$	$O_2 \rightarrow$	$2 H_2O$
Before	10.0 g H_2 	1.0 g O_2 	0.00 g H_2O
Change 10.0 g H_2 1.0 g O_2 0.00 g H_2O 0.00 g H_2O	-10.0 g H_2 -1.0 g O_2 0.00 g H_2O 	+1.0 g O_2 +1.0 g O_2 -1.0 g O_2	+10.0 g H_2O +10.0 g H_2O -10.0 g H_2O
After	0.00 g H_2 	0.00 g O_2	10.0 g H_2O produced

Change in the limiting reactant



- b. What mass of water can be produced? Show your work.

$$0.031 \text{ mol } O_2 \left(\frac{2 \text{ mol } H_2O}{1 \text{ mol } O_2} \right) = 0.062 \text{ mol } H_2O$$

- c. Which reactant is present in excess, and what mass of that reactant remains after the reaction is complete?

H_2 is present in excess
initial mass of H_2 remains

$$10.0 \text{ g } H_2 \left(\frac{2.016 \text{ g}}{1 \text{ mol}} \right) = 4.96 \text{ g } H_2 \text{ remains}$$

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DJ Losen



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