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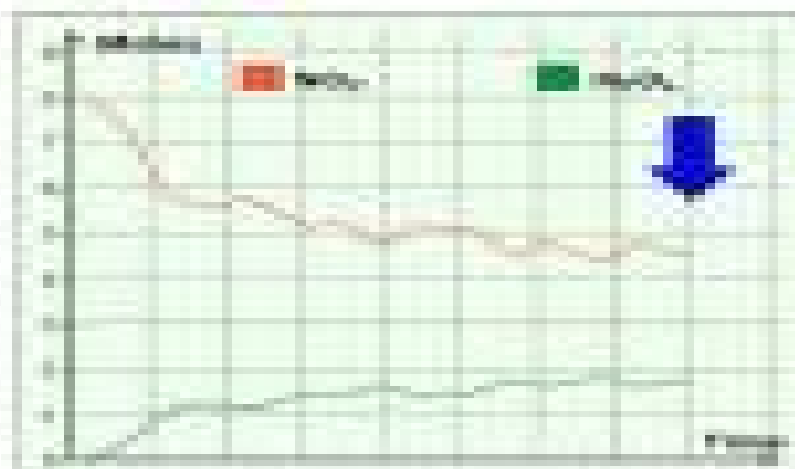
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### Equilibrium (Le Chatelier's)

A closed flask containing hydrogen gas at high pressure (100 atm) and 500 K is cooled. The flask contains the following chemical equilibrium mixture. What is most likely to happen to the flask's pressure?



- The concentration of the hydrogen gas will increase.
- The reaction mixture will shift to the right to produce more hydrogen gas.
- The rate of the forward reaction will increase.
- The reaction will shift to the right to produce more hydrogen gas.

**Correct Answer: B** The flask is at equilibrium. If the temperature is lowered, the system will shift to the right to produce more hydrogen gas.

Hydrogen gas is at high pressure (100 atm) and 500 K. The flask contains the following chemical equilibrium mixture. What is most likely to happen to the flask's pressure?

# Explorelearning Equilibrium And Concentration Gizmo

**Clifford Lane**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

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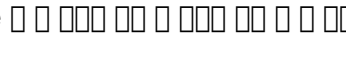
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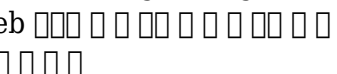

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