

GROUP #: OR I worked on this part alone , alone

Equilibrium Exploration 2: Equilibrium and Pressure

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.



Vocabulary: Dalton's law, Le Châtelier's principle, partial pressure, pressure

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

A typical scuba tank has a volume of 11 liters and can support a diver for one hour. An adult breathes about 3 liters of air with each breath.

1. How can an 11-liter tank give a diver enough oxygen for one hour?

The amount of oxygen in the tank is very dense.

2. Why are diving cylinders made of thick, reinforced aluminum or steel?

To prevent the cylinder from being broken by the immense pressure within it.

Gizmo Warm-up

Gases consist of billions of tiny particles in constant motion, colliding with each other and the walls of the container. The sum of all these collisions creates **pressure** on the walls of the container. In theory, any amount of gas can be squeezed into a container if the container is strong enough to withstand the gas pressure.



The *Equilibrium and Pressure* Gizmo shows a mixture of gases in chamber. The lid of the chamber can move up or down.

1. Check that **Reaction 1** is selected. Use the **Moles NO₂** slider to increase the number of NO₂ molecules in the chamber. How does this affect the volume of the chamber?

It increases the volume of the chamber.

2. Notice the weights on the lower right side of the Gizmo. Drag several of these weights to the lid. How does this affect the volume of the chamber?

As you add more weights, it decreases the volume of the chamber.

Explore Learning Equilibrium And Pressure Gizmo Answers

J Elliott



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