

Student Exploration: Solubility and Temperature

Summary: Key Points

Solubility refers to the ability of a substance to dissolve in a solvent. It is measured in terms of the amount of solute that can dissolve in a given amount of solvent at a specific temperature. Solubility is affected by temperature, pressure, and the nature of the solute and solvent. For example, the solubility of most solids increases with temperature, while the solubility of gases decreases with temperature. The solubility of a substance is also affected by the nature of the solute and solvent. For example, ionic compounds are more soluble in polar solvents like water, while non-polar compounds are more soluble in non-polar solvents like oil.

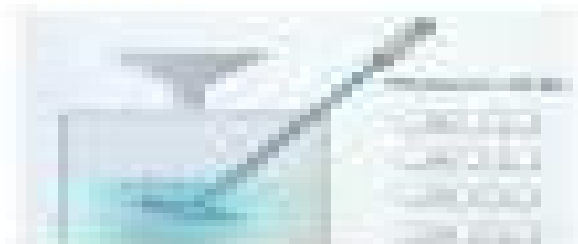
Key Concepts: Solubility and Temperature

1. What happens when you add a substance to a solvent? (The substance dissolves.)
2. What happens when you add a substance to a solvent at a different temperature? (The substance dissolves more or less.)
3. What happens when you add a substance to a solvent at a different pressure? (The substance dissolves more or less.)
4. What happens when you add a substance to a solvent at a different temperature and pressure? (The substance dissolves more or less.)

Key Concepts: Solubility and Temperature

At different temperatures, the solubility of a substance is different. For example, the solubility of a substance is higher at a higher temperature than at a lower temperature. This is because the molecules of the substance have more energy at a higher temperature, which allows them to move more freely and dissolve more easily in the solvent.

However, the solubility of a substance is also affected by the nature of the solute and solvent. For example, ionic compounds are more soluble in polar solvents like water, while non-polar compounds are more soluble in non-polar solvents like oil.



1. In this solution, what is the solute? (The substance being dissolved.) What is the solvent? (The substance being dissolved in.)
2. Add 10 g of solute to 100 mL of solvent. What happens? (The solute dissolves.)
3. Add 20 g of solute to 100 mL of solvent. What happens? (The solute dissolves more.)
4. Add 30 g of solute to 100 mL of solvent. What happens? (The solute dissolves even more.)

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