

Unit 7 Worksheet 1: Molar-mass calculations

Four marks will be given for each question.

1. Write down balanced chemical equations

a) for the reaction of propanoic acid, acetic acid and acetic acid with calcium hydroxide

b) for the reaction of calcium hydroxide with calcium hydroxide

c) Please call up hydrogen in (a)

2. Hydrogen sulfide gas, which smells like rotten eggs, burns in air to produce water and sulfur. How many moles of oxygen are needed to completely burn 10 moles of hydrogen sulfide?

Equation: _____

Answer: _____

Change: _____

Notes: _____

3. Propane, C_3H_8 , burns in air to form carbon dioxide and water. If 10 moles of carbon dioxide are formed, how many moles of propane were burned?

Equation: _____

Answer: _____

Change: _____

Notes: _____

4. Acetylene, C_2H_2 , has been used in welding. Hydrogen and oxygen gases are both diatomic and react to form acetylene. How many moles of acetylene are formed from 10 moles of hydrogen gas?

Equation: _____

Answer: _____

Change: _____

Notes: _____

Modeling Chemistry U8 Test

RC Schank



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