

- Solution C1, an aqueous solution of potassium iodate (V) of concentration 4.06 g/dm^3 .
- Solution C2, an aqueous solution of sodium thiosulphate of unknown concentration.
- Solution C3, an acidified solution of potassium iodide.
- Solution C4, starch indicator solution.

Section 1

You are required to:

- React solution C3 with C1 to liberate a certain quantity of iodine.
- Titrate the iodine liberated in the reaction between C3 and C1 against solution C2 to determine the molar concentration of C2.

Procedure

- Fill the burette with solution C2.
- Pipette 25.0 cm^3 of solution C1 into a 250 ml conical flask.
- Using a measuring cylinder, transfer 15.0 cm^3 of solution C3 into a conical flask with solution C1.
- Titrate the solution in the conical flask against solution C2 from the burette until the brown colour just changes to pale yellow.
- Using a 10 ml measuring cylinder, add 5 cm^3 of solution C4 into the mixture in the conical flask and then continue titrating until the blue colour just disappears.
- Record your titration results in table 1 below.
- Repeat the procedure (b) to (f) above two more times and complete table 1 below.

(4 marks)

Table 1	I	II	III
Final burette reading (cm^3)			
Initial burette reading (cm^3)			
Volume of solution C2 (cm^3) added			

- (8) Determine the average volume of C2 used.

(1 mark)

- (i) Potassium iodate (V) solution reacts with acidified potassium iodide solution to liberate iodine as per the following ionic equation.



- (ii) Sodium thiosulphate reduces iodine to iodide ions as per the ionic equation below:



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