

## Chapter 1 - Introduction

**1.1** It is given at 20°C, water has condensed  $10^{-3}$  g molecules/liter. How many  $\text{H}_2\text{O}^+$  molecules per  $\text{cm}^3$  of atmosphere's number is  $6.02 \times 10^{23}$  molecules per mole, what air pressure does this represent?

**Solution:** The mass of one molecule of air may be approximated as

$$m = \frac{\text{Molecular weight}}{\text{Avogadro's number}} = \frac{(28.97 \text{ g/mol})}{6.02 \times 10^{23} \text{ molecules/mol}} = 4.815 \times 10^{-26} \text{ g}$$

Then the density of air containing  $10^{-3}$  molecules per  $\text{cm}^3$  is, in SI units,

$$\begin{aligned} \rho &= \left( 10^{-3} \frac{\text{molecules}}{\text{cm}^3} \right) \left( 4.815 \times 10^{-26} \frac{\text{g}}{\text{molecule}} \right) \\ &= 4.815 \times 10^{-29} \frac{\text{g}}{\text{cm}^3} = 4.815 \times 10^{-30} \frac{\text{kg}}{\text{m}^3} \end{aligned}$$

Finally, from the perfect gas law, Eq. (1.10), at  $20^\circ\text{C} = 293 \text{ K}$ , we obtain the pressure

$$p = \rho RT = \left( 4.815 \times 10^{-30} \frac{\text{kg}}{\text{m}^3} \right) \left( 293 \frac{\text{K}}{\text{K}} \right) \left( 287 \frac{\text{J}}{\text{kg} \cdot \text{K}} \right) = 4.04 \times 10^{-26} \text{ Pa}$$

**1.2** The earth's atmosphere can be modeled as a uniform layer of air of thickness 10 km and average density that equals that of air at 10 km. How does this help to estimate the total mass and total number of molecules of air in the entire atmosphere of the earth?

**Solution:** Let  $R_E$  be the earth's radius = 6375 km. Then the total mass of air in the atmosphere is

$$\begin{aligned} m_A &= \int_V \rho \, dV = \rho_{10 \text{ km}} (4\pi R_E^2) (10 \text{ km}) = \rho_{10 \text{ km}} (4\pi R_E^2) (\text{air thickness}) \\ &= (1.25 \text{ kg/m}^3) (4\pi (6375 \text{ km})^2) (10 \text{ km}) = 1.02 \times 10^{22} \text{ kg} \quad \text{Ans.} \end{aligned}$$

Knowing by the mass of one molecule =  $4.815 \times 10^{-26} \text{ g}$  from Prob. 1.1, already we obtain the total number of molecules in the earth's atmosphere

$$N_{\text{molecules}} = \frac{(\text{atmosphere mass})}{(\text{mass molecule})} = \frac{1.02 \times 10^{22} \text{ kg}}{4.815 \times 10^{-26} \text{ g/molecule}} = 2.12 \times 10^{48} \text{ molecules} \quad \text{Ans.}$$

# Fluid Mechanics White Solutions Manual 5th Edition

**Minjie Lin**



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