

Chemistry Unit 8 Worksheet 4

Samples of Every Kind of Problem

Let us suppose that our "paper" writer is a chemist who has to write out the products of the reaction. Show the products and label the phase. The water is liquid (not "drip" water).

1. Chlorine gas reacts with a metal to produce metal chloride. BaCl_2 and Ba react. Write the balanced equation for products potassium chloride. KCl is a solid. Is metal "reactant" gas?

Reactants	BaCl_2	Ba	BaCl_2	Ba
Products	BaCl_2	Ba	BaCl_2	Ba
Amount	1.00	1.00	1.00	1.00
Mass	208	137	208	137

Reaction
balanced

2. Silver metal is displaced from a solution, copper metal reacts with hydrochloric acid to produce copper(II) chloride and liquid water. Write the balanced equation for hydrochloric acid to react with silver metal. Write the mass of the product.

Reactants	HCl	Ag	HCl	Ag
Products	HCl	Ag	HCl	Ag
Amount	1.00	1.00	1.00	1.00
Mass	36	108	36	108

Reaction
balanced

3. Chlorine, Cl_2 , reacts with a metal to produce metal chloride. Chlorine gas reacts with silver. Write the balanced equation for silver to react with chlorine gas. Write the mass of the product.

Reactants	Cl_2	Ag	Cl_2	Ag
Products	Cl_2	Ag	Cl_2	Ag
Amount	1.00	1.00	1.00	1.00
Mass	71	108	71	108

Reaction balanced
mass = 179 g

4. Chlorine gas reacts with a metal to produce metal chloride. Chlorine gas reacts with silver. Write the balanced equation for silver to react with chlorine gas. Write the mass of the product.

$\text{Cl}_2 + 2\text{Ag} \rightarrow 2\text{AgCl}$ mass = 236 g

Reactants	Cl_2	Ag	Cl_2	Ag
Products	Cl_2	Ag	Cl_2	Ag
Amount	1.00	1.00	1.00	1.00
Mass	71	108	71	108

Reaction balanced
mass = 236 g

Modeling Chemistry Unit 8 1 Answers

Judd E. Hollander



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