

Mixed Gas Laws Worksheet

- 1) How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K?

$$V = 98 \text{ L}$$

$$P = 2.8 \text{ atm}$$

$$T = 292 \text{ K}$$

$$n = ?$$

$$R = 0.082 \frac{\text{L atm}}{\text{mol K}}$$

$$PV = nRT$$

$$n = \frac{PV}{RT}$$

$$= \frac{2.8 \times 98}{0.082 \times 292} = \underline{\underline{11.46 \text{ mole}}}$$

- 2) If 5.0 moles of O_2 and 3.0 moles of N_2 are placed in a 30.0 L tank at a temperature of 25°C , what will the pressure of the resulting mixture of gases be?

$$5 \text{ moles } \text{O}_2$$

$$3 \text{ moles } \text{N}_2$$

$$V = 30 \text{ L}$$

$$T = 25^\circ\text{C} \rightarrow 273 + 25$$

$$P = ?$$

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{8 \times 0.082 \times 298}{30 \text{ L}}$$

$$= \underline{\underline{6.5 \text{ atm}}}$$

- 3) A balloon is filled with 35.0 L of helium in the morning when the temperature is 20.0°C . By noon the temperature has risen to 45.0°C . What is the new volume of the balloon?

$$① V_1 = 35 \text{ L}$$

$$① T_1 = 20^\circ\text{C} = 293 \text{ K}$$

$$② T_2 = 45^\circ\text{C} = 318 \text{ K}$$

$$② V_2 = ?$$

Charles law

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$$\frac{35}{20} = \frac{V_2}{45}$$

$$V_2 = \frac{35 \times 318}{293} = \underline{\underline{37.78 \text{ L}}}$$

- 4) A 35 L tank of oxygen is at 315 K with an internal pressure of 190 atmospheres. How many moles of gas does the tank contain?

$$V = 35 \text{ L}$$

$$T = 315 \text{ K}$$

$$P = 190 \text{ atm}$$

$$n = ?$$

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{190 \times 35}{0.082 \times 315}$$

$$= \underline{\underline{257.45 \text{ mol}}}$$

Mixed Gas Law Worksheet Answers

Jicheng Xie

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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the needs or interests of your student materials for each semester are independent of one another to allow flexibility

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Table of Contents Mixed Gas Law Worksheet Answers

1. Understanding the eBook Mixed Gas Law Worksheet Answers
 - The Rise of Digital Reading Mixed Gas Law Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Mixed Gas Law Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mixed Gas Law Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mixed Gas Law Worksheet Answers
 - Personalized Recommendations
 - Mixed Gas Law Worksheet Answers User Reviews and Ratings
 - Mixed Gas Law Worksheet Answers and Bestseller Lists
5. Accessing Mixed Gas Law Worksheet Answers Free and Paid eBooks
 - Mixed Gas Law Worksheet Answers Public Domain eBooks
 - Mixed Gas Law Worksheet Answers eBook Subscription Services
 - Mixed Gas Law Worksheet Answers Budget-Friendly Options

6. Navigating Mixed Gas Law Worksheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Mixed Gas Law Worksheet Answers Compatibility with Devices
 - Mixed Gas Law Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mixed Gas Law Worksheet Answers
 - Highlighting and Note-Taking Mixed Gas Law Worksheet Answers
 - Interactive Elements Mixed Gas Law Worksheet Answers
8. Staying Engaged with Mixed Gas Law Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mixed Gas Law Worksheet Answers
9. Balancing eBooks and Physical Books Mixed Gas Law Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mixed Gas Law Worksheet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mixed Gas Law Worksheet Answers
 - Setting Reading Goals Mixed Gas Law Worksheet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mixed Gas Law Worksheet Answers
 - Fact-Checking eBook Content of Mixed Gas Law Worksheet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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