

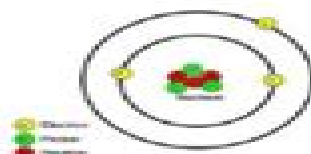
## Atomic Structure – Study Guide

### Grade 8 – Physical Science

**Atom** -- smallest particle of an element that has all the properties of that element.

**3 main subatomic particles:**

1. protons
2. neutrons
3. electrons



**nucleus** -- core of the atom, it is positively charged 99.9 % of the atoms mass is in the nucleus. Protons and neutrons are found in the **nucleus**.

**1. Protons** -- have a positive (+) charge. Found in the nucleus of the atom.

The **number of protons** determines what element an atom is.

**Example:** Carbon has 6 protons

**Atomic Number** = the **number of protons**, each element has it own unique number of protons.

**2. Neutrons** -- a neutral charge (no charge) found in the nucleus. These determine whether an element is radioactive.

Protons and neutrons are about *one atomic mass unit (amu)*. Electrons have a much smaller mass -- it takes almost 2000 electrons to equal 1 amu.

**Atomic Mass** = the total number of protons and neutrons.

**Mass Number** = Atomic Mass that is rounded.

To find just how many **neutrons** an atom has:

# neutrons = atomic mass – atomic #

**Isotopes** - atoms of the same element that have different number of neutrons.

**Carbon-12**

6 protons

6 neutrons

**Carbon-14**

6 protons

8 neutrons

**NOTICE** – the number of **protons DOES NOT** change if it is to be carbon -- only the number of neutrons.

**3. Electrons** - have a negative (-) charge. Found moving around the atom is a space called orbital's (shells).

Those electrons farthest away from the nucleus are called **valence electrons**, and are involved in the formation of chemical bonds.

In an **uncharged atom** the **number of protons is equal to the number of electrons**.

**ion** - when an atom gains or loses an *electron*. (has a charge)

If an atom gains electrons = negative (-) charge

If an atom loses electrons = positive (+) charge

**How to find:**

# protons = atomic #

# neutrons = atomic mass – atomic #

# electrons = atomic # – charge (if any)

# Matter And Atomic Structure Study Guide Answers

**RD Boyd**



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## **Table of Contents Matter And Atomic Structure Study Guide Answers**

1. Understanding the eBook Matter And Atomic Structure Study Guide Answers
  - The Rise of Digital Reading Matter And Atomic Structure Study Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Matter And Atomic Structure Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Matter And Atomic Structure Study Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Matter And Atomic Structure Study Guide Answers

- Personalized Recommendations
  - Matter And Atomic Structure Study Guide Answers User Reviews and Ratings
  - Matter And Atomic Structure Study Guide Answers and Bestseller Lists
5. Accessing Matter And Atomic Structure Study Guide Answers Free and Paid eBooks
    - Matter And Atomic Structure Study Guide Answers Public Domain eBooks
    - Matter And Atomic Structure Study Guide Answers eBook Subscription Services
    - Matter And Atomic Structure Study Guide Answers Budget-Friendly Options
  6. Navigating Matter And Atomic Structure Study Guide Answers eBook Formats
    - ePub, PDF, MOBI, and More
    - Matter And Atomic Structure Study Guide Answers Compatibility with Devices
    - Matter And Atomic Structure Study Guide Answers Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Matter And Atomic Structure Study Guide Answers
    - Highlighting and Note-Taking Matter And Atomic Structure Study Guide Answers
    - Interactive Elements Matter And Atomic Structure Study Guide Answers
  8. Staying Engaged with Matter And Atomic Structure Study Guide Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Matter And Atomic Structure Study Guide Answers
  9. Balancing eBooks and Physical Books Matter And Atomic Structure Study Guide Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Matter And Atomic Structure Study Guide Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Matter And Atomic Structure Study Guide Answers
    - Setting Reading Goals Matter And Atomic Structure Study Guide Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Matter And Atomic Structure Study Guide Answers

- Fact-Checking eBook Content of Matter And Atomic Structure Study Guide Answers
- Distinguishing Credible Sources

### 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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