

Name: _____

Genetics Practice Problems (Remote Edition)



1. For each genotype below, indicate whether it is heterozygous (**He**) or homozygous (**Ho**)

AA **Ho** Ee **He** Bb **He** ff **Ho**
Pp **He** gg **Ho** Dd **He** HH **Ho**

2. For each of the **genotypes** below determine what **phenotypes** would be possible.

Purple flowers are dominant to white

PP Purple
Pp Purple
pp White

Brown eyes are dominant to blue

BB Brown
Bb Brown
bb Blue

3. For each **phenotype** below, list the **genotypes** (remember to use the letter of the dominant trait)

Straight hair is dominant to curly

SS straight
Ss straight
ss curly

Tail spikes are dominant to plain tails

SS spikes
Ss spikes
ss plain



4. Complete the Punnett squares for each of the crosses listed below.

Tall (T) plants are dominant to short (t)

Tt x tt

	<u>T</u>	<u>t</u>
<u>t</u>	Tt	tt
<u>t</u>	Tt	tt

What percentage of the offspring are tall? **50%** short? **50%**

Tt x Tt

	<u>T</u>	<u>t</u>
<u>T</u>	TT	Tt
<u>t</u>	Tt	tt

What percentage of the offspring are tall? **75%** short? **25%**



5. A **homozygous tall** plant is crossed with a **short** plant.

Genetics Practice Problems Answer Key Sutherland Ffa

TD Snyder



Genetics Practice Problems Answer Key Sutherland Ffa:

The book delves into Genetics Practice Problems Answer Key Sutherland Ffa. Genetics Practice Problems Answer Key Sutherland Ffa is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Genetics Practice Problems Answer Key Sutherland Ffa, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:
 - Chapter 1: Introduction to Genetics Practice Problems Answer Key Sutherland Ffa
 - Chapter 2: Essential Elements of Genetics Practice Problems Answer Key Sutherland Ffa
 - Chapter 3: Genetics Practice Problems Answer Key Sutherland Ffa in Everyday Life
 - Chapter 4: Genetics Practice Problems Answer Key Sutherland Ffa in Specific Contexts
 - Chapter 5: Conclusion
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 4. In chapter 3, the author will examine the practical applications of Genetics Practice Problems Answer Key Sutherland Ffa in daily life. The third chapter will showcase real-world examples of how Genetics Practice Problems Answer Key Sutherland Ffa can be effectively utilized in everyday scenarios.
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 6. In chapter 5, this book will draw a conclusion about Genetics Practice Problems Answer Key Sutherland Ffa. This chapter will summarize the key points that have been discussed throughout the book.
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the balanced equation shown below if the reaction of 90.6 grams of CO produces 36.7 grams of C₃H₈ what is the percent yield
 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation
 $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$
8.3 mole H₂S $\times$ 18g H₂O / 1 mole H₂O = 149.4g H₂O is your theoretical yield so percent yield = $\frac{137.1\text{g H}_2\text{O}}{149.4\text{g H}_2\text{O}} \times 100 = 91.77\%$
problems 6 & 8 are similar to problem 5

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 $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
what is the limiting reagent if 8.52 g are actually formed what is the

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 $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$
if the reaction of 21.3 grams of O₂ produces a 47.3% yield how many grams of CO would be produced
step one figure out how many moles of the limiting reagent you have
 $\frac{21.3 \text{ g O}_2}{32 \text{ g/mol}} = 0.666 \text{ moles O}_2$

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web percent yield is a measure of the efficiency of the experimental design
yield = $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$
in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction
 $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
4 Al + 3 O₂ $\rightarrow$ 2 Al₂O₃
b determine the theoretical yield of Al₂O₃
 $\frac{160.0 \text{ g Al}}{27 \text{ g/mol}} \times \frac{1 \text{ mol Al}_2\text{O}_3}{2 \text{ mol Al}} \times 101.96 \text{ g/mol} = 302.3 \text{ g Al}_2\text{O}_3$
c determine the percent yield
 $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$

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web 7 Na₂C₂O₄ is the limiting reactant percent yield = 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H₂ are left over 13

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