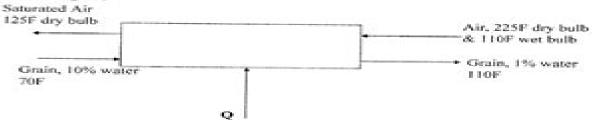
- Air at 45°C and 10% relative humidity is to be humidified adiabatically to 60% relative humidity. Use the psychometric chart to obtain the following numbers:
 - (a) (5 points) The dry bulb and adiabatic saturation temperature of the final air.
 - (b) (5 points) The dew point of the final air
 - (c) (5 points) The absolute humidity and humid volume of final air.
 - (d) (5 points) The rate at which water must be added to humidify 15m3/min of the entering air.
- 2. (20 points) Two aqueous sulfuric acid solutions containing 20.0 wt% H2SO4 (specific gravity=1.139) and 60.0 wt % H₂SO₄ (specific gravity=1.498) are mixed to form a 4.00 molar solution (specific gravity=1.213). Assuming 100 kg of 20% solution as the basis calculate:
 - (a) (10 points) the mass fraction of sulfuric acid in the product solution.
 - (b) (10 points) the feed ratio (liters of 20% solution/liter of 60% solution).
- 3. (30 points) A rotary dryer operating at atmospheric pressure dries 10 tons/day of wet grain at 70F, from a moisture content of 10% to 1% moisture. The air flow is countercurrent to the flow of grain, enters at 225F dry bulb and 110F wet-bulb temperature, and leaves at 125F dry-bulb (see figure).



Determine: (a) the humidity of entering and leaving air if the latter is saturated, (b) the water removal in pounds per hour. (c) the daily product output in pounds per day, and (d) the heat input to the dryer. Assume that there is no heat loss from the dryer, that the grain is discharged at 110F, and that its specific heat is 0.18.

- 4. (30 points) A fuel gas containing 85 mole % methane and the balance ethane is burned completely with pure oxygen at 25 °C at latm. Suppose the reactor is continuous. Take a basis of calculation of 1 mol/s of the fuel gas, combusted with stoichiometric amount of oxygen. (a) (10 points) Please calculate the rate at which heat must be transferred from the reactor if the
 - products are brought back down to 25 °C.
 - (b) (15 points) Please calculate the final temperature of the combustion gas if the process is under adiabatic condition at constant pressure 1atm.
 - (c) (5 points) Please calculate the final volume of the combustion gas if the process is under adiabatic condition at constant pressure latm. If the volume of the reactor can only be constant and the pressure can change, what do you expect to happen to the reactor?

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following sections 1 ER in stress signalling and adaptation 2 ER structure and biosynthetic functions 3 Regulation of protein processing 4 Regulation of programmed cell death Government Reports Announcements & Index ,1979 Harper's Weekly John Bonner, George William Curtis, Henry Mills Alden, Samuel Stillman Conant, Montgomery Schuyler, John Foord, Richard Harding Davis, Carl Schurz, Henry Loomis Nelson, John Kendrick Bangs, George Brinton McClellan Harvey, Norman Hapgood, 1875 Text Book of Life Science Rajveer Singh Chauhan, 2010 Life Science: Origins and Scientific Theory Parent Lesson Planner Master Books, 2013-08-01 This Life Science Origins Scientific Theory contains materials for use with Evolution The Grand Experiment Vol 1 and Living Fossils Course Description Upon completion of this course students will have a thorough understanding of the theory of evolution and its limits Students will develop scientific critical thinking skills through careful analysis of evidence and comparing the merits of different theories Students will study paleontology biology and geology as they relate to the study of origins through an exploration of living fossils Workflow Step 1 Teacher leads Discussion Questions Step 2 Student watches DVD if applicable Step 3 Student reads chapter Step 4 Student is given and completes Chapter Objectives after reading the chapter Step 5 Teacher administers Chapter Test Step 6 Teacher administers Sectional Exams where indicated Step 7 Teacher administers Comprehensive Final Exams where indicated

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