Use the information in these tables to complete the graph as described below

Symbol	Atomic Radius (Picometers)	First Ionization energy (kilojoules/mole)	Electronegativity (4-point scale)
H	31	1312	2.1
IL-i	128	520	1.0
Nia	166	496	0.9
K	203	410	0.8
Rb	220	403	0.8
Cs	244	376	0.7

Symbol	Atomic Radius (Picometers)	First lonization energy (kilojoules/mole)	Electronegativity (4-point scale)
Na	166	496	0.9
Mg	141	738	1.2
AL	121	578	1.5
Sii	111	787	1.8
P	107	1012	2.1
S	105	1000	2.5
CI	102	1251	3.0
Ar	106	1521	

- Using colored pencils or pen, list under the symbol (in this order!) the Atomic Radius, First Ionization Energy, and Electronegativity.
- Use a different color for each property. Example: Write all of the Atomic radius values in red, all of the First
 ionization energies in green, and all of the Electronegativities in blue. Of course, you can pick any colors that are
 available, as long as you are consistent.
- 3. Observe the trends in each property as you go down the Alkali metal group, and as you go across Period 3.
- Write out each of the statements written on the opposite side, completing each statement with the observed trend (increase or decrease)

Li	Across	a period (-	-) electrone	gativity ter	nds to		
Na 166 496 0.9	Mg	Al	Si	Р	s	CI	A
К	Down	a Group (‡) atomic rac	lius tends t	o		
Rb	01,000,000,000) first ioniza) electroneç		_		

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