

The Ideal and Combined Gas Laws $PV = nRT$ or $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$

Use your knowledge of the ideal and combined gas laws to solve the following problems. If it involves moles or grams, it must be $PV = nRT$

- 1) If four moles of a gas at a pressure of 5.4 atmospheres have a volume of 120 liters, what is the temperature?

$$(5.4)(120) = 4(.0821)T$$

$$T = 1973.2 \text{ K}$$

- 2) If I initially have a gas with a pressure of 84 kPa and a temperature of 35° C and I heat it an additional 230 degrees, what will the new pressure be? Assume the volume of the container is constant.

$$\frac{84}{(35+273)} = \frac{P_2}{(35+230+273)} \quad P_2 = 146.7 \text{ kPa}$$

- 3) My car has an internal volume of 2600 liters. If the sun heats my car from a temperature of 20° C to a temperature of 55° C, what will the pressure inside my car be? Assume the pressure was initially 760 mm Hg.

$$\frac{760 \text{ mmHg}}{(20+273)} = \frac{P_2}{(55+273)} \quad P_2 = 850 \text{ mmHg}$$

- 4) How many moles of gas are in my car in problem #3?

$$\frac{850 \text{ mmHg}}{760 \text{ mmHg}} = 1.13 \text{ atm}$$

$$2600(1.13) = n(.0821)(55+273)$$

$$n = 107.1 \text{ moles}$$

- 5) A toy balloon filled with air has an internal pressure of 1.25 atm and a volume of 2.50 L. If I take the balloon to the bottom of the ocean where the pressure is 95 atmospheres, what will the new volume of the balloon be? How many moles of gas does the balloon hold? (Assume $T = 285 \text{ K}$)

$$1.25(2.50) = 95(V_2)$$

$$V_2 = 0.032 \text{ L}$$

$$95(.032) = n(285)(.0821)$$

$$n = 0.1299 \text{ moles}$$

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