

Gas Laws Worksheet

Overview/Introduction

In this assignment, you will perform calculations using the various gas laws. This will help you understand the relationship between pressure, volume, and temperature in the gaseous state.

Directions

Complete the problems for each section. If you used steps to complete the calculations or determine the answer, please include them with your answer. All answers and work shown should be neat, clear, and legible.

Equations

- ☐ $1 \text{ atm} = 760.0 \text{ mmHg} = 101.3 \text{ kPa} = 14.7 \text{ lbs/in}^2$
 - ☐ $1 \text{ mole gas} = 22.4 \text{ L at STP}$
1. If 34.7 L of nitrogen at 748 mmHg are compressed to 725 mmHg at constant temperature, what is the new volume of nitrogen?
 2. A gas with a volume of 5.0 L at a pressure of 210 kPa is allowed to expand to a volume of 12.0 L. What is the pressure in the container if the temperature remains constant?
 3. A gas occupies 700.0 mL at a temperature of 27.0 °C. What is the volume of gas at 120.0 °C assuming pressure did not change?
 4. A container of 5.00 L of gas is collected at 93 K and then allowed to expand to 25.0 L. What must the new temperature be assuming pressure is constant?
 5. On a 26 °C summer day, the pressure in Tony Cousteau's backup oxygen tank read 900 mmHg before he dove into the ocean. He dove to a depth where the temperature was -23 °C. What is the pressure in Jacques's backup tank at that temperature?
 6. A paint spray can has a pressure of 30 lbs/in² at a temperature of 27 °C. If the gasses in the can reach a pressure of 90 lbs/in², the can will explode. What temperature must the gasses in the spray can be at in order for the spray can to explode? (Assume constant volume)
 7. What is the density of carbon monoxide at STP?
 8. 50 g of nitrogen gas (N₂) has a volume of ____ liters at STP.

Gas Laws Practice Calculations Answers Unit Lesson 01

Ying-Ying Zheng

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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