

Gas Laws Problem Packet

Gas Laws Packet

Name _____ Date _____ Period _____

Introduction

Standard Conditions

Unit	Standard Pressure
atm	1.013
mmHg	760
cmHg	76
torr	760

22.4 L = 1 mol at STP

• STP (standard temperature and pressure) is 0°C and 1 atm.

• The gas laws:

Gas Law	Relationship
Boyle's Law	$P_1V_1 = P_2V_2$
Charles's Law	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$
Gay-Lussac's Law	$\frac{P_1}{T_1} = \frac{P_2}{T_2}$
Avogadro's Law	$\frac{V_1}{n_1} = \frac{V_2}{n_2}$
Combined Gas Law	$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$
Ideal Gas Law	$PV = nRT$

• The constant R used in the ideal gas law depends on the unit of pressure.

a. $R = 0.0821 \frac{\text{atm} \cdot \text{L}}{\text{mol} \cdot \text{K}}$

b. $R = 8.314 \frac{\text{J}}{\text{mol} \cdot \text{K}}$

Notes:

- Temperature must always be in Kelvin ($K = C + 273$).
- If you are given two different units of pressure, convert!
- Create two columns below (initial and final). Write what you're given in the problem. Mark the two that use the variable you're given, and solve for the unknown. For ideal gas laws, write what you are given, insert those in equation and solve for unknown.
- Avogadro's Law: Solve for n (moles) to use ideal gas law.
- Graham's Law: You will need to know the molecular mass for these problems. Double and triple check you have the right of effusion and molecular masses in the correct units before solving the equation (this is an $\text{m}^{-1/2}$ relationship). You must evaluate the square root or square both sides of the equation before moving things.



All Gas Laws Included!



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