

MATH SKILLS TRANSPARENCY WORKSHEET

16

Using Mole Ratios

 Use with Chapter 11,
Section 11.2

For each of the following problems, write the balanced chemical equation that represents the reaction. Then complete the table below by identifying the known substance, the unknown substance, and the mole ratio that you would use to solve each problem correctly.

1. Copper(II) oxide (CuO) decomposes into copper (Cu) and oxygen (O_2) gas. What mass of copper will be produced by the decomposition of 1.25 kg CuO ?



2. Ammonia (NH_3) is produced by the reaction of nitrogen (N_2) and hydrogen (H_2) gases. How much ammonia will be produced if 22.8 g H_2 reacts with excess N_2 ?



3. The reaction of sodium (Na) and water (H_2O) produces sodium hydroxide (NaOH) and hydrogen (H_2) gas. What mass of hydrogen gas is produced if 17.54 g NaOH is produced by the reaction?



4. The combustion of acetic acid ($\text{HC}_2\text{H}_3\text{O}_2$) produces carbon dioxide (CO_2) and water (H_2O). What mass of carbon dioxide will be produced from the combustion of 25.0 g $\text{HC}_2\text{H}_3\text{O}_2$?



5. 20.2 g of iron(III) sulfide (Fe_2S_3) was prepared by heating iron (Fe) and excess sulfur (S). What mass of iron was used in the preparation?



Problem	Chemical Formula of Known Substance	Chemical Formula of Unknown Substance	Mole Ratio
1	CuO	Cu	$\frac{2 \text{ mol Cu}}{2 \text{ mol CuO}}$
2	H_2	NH_3	$\frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3}$
3	NaOH	H_2	$\frac{2 \text{ mol NaOH}}{1 \text{ mol H}_2}$
4	$\text{HC}_2\text{H}_3\text{O}_2$	CO_2	$\frac{1 \text{ mol HC}_2\text{H}_3\text{O}_2}{2 \text{ mol CO}_2}$
5	Fe_2S_3	Fe	$\frac{1 \text{ mol Fe}_2\text{S}_3}{2 \text{ mol Fe}}$

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