

1.2.1 Compare (qualitatively) the relative strengths of ionic, covalent, and metallic bonds.

1. Compare the relative strengths of the three different types of bonds.

ionic > covalent > metallic

2. Which statement compares the amount of energy needed to break the bonds in CaCl_2 (E1) and $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ (E2)?

- a. $\text{E1} > \text{E2}$, as CaCl_2 is a covalent compound.
- b. $\text{E1} < \text{E2}$, as CaCl_2 is a covalent compound.
- ☒ c. $\text{E1} > \text{E2}$, as CaCl_2 is an ionic compound.
- d. $\text{E1} < \text{E2}$, as CaCl_2 is an ionic compound.

1.2.2 Infer the type of bond and chemical formula formed between atoms.

1. Which pair of elements would *most likely* bond to form a covalently bonded compound?

- a. sodium and fluorine
- b. barium and chlorine
- ☒ c. phosphorus and oxygen
- d. magnesium and sulfur

2. For each pair of atoms, predict whether the bond formed between the atoms is either ionic or covalent, and write the formula for the predicted compound.

- 1. Na and O: ionic, Na_2O
- 2. S and F: covalent, SF_2
- 3. Ag and N: ionic, Ag_3N
- 4. Te and H: covalent, TeH_2

ionic = M + NM (or polyatomic ion)
covalent = all NM
metallic = all M

3. Which statement describes the compound formed between sodium and oxygen?

- a. It is NaO_2 , which is ionic.
- b. It is NaO_2 , which is covalent.
- ☒ c. It is Na_2O , which is ionic.
- d. It is Na_2O , which is covalent.

1.2.3 Compare inter- and intra- particle forces.

1. Rank the following substances in the order in which they would evaporate, justifying the order of placement for each (using those inter- and intra- particle forces).

- 1. Water (H_2O)
- 2. Methane (CH_4)
- 3. Sodium chloride (NaCl)
- 4. Phosphorus trifluoride (PF_3)

1. CH_4 - LDFs (weakest)
2. PF_3 - LDFs, dipole-dipole
3. H_2O - LDFs, dipole-dipole, H-bonding
4. NaCl - ionic bond (strongest forces)

2. At STP, fluorine is a gas and iodine is a solid. Why?

- a. Fluorine has lower average kinetic energy than iodine.
- b. Fluorine has higher average kinetic energy than iodine.
- ☒ c. Fluorine has weaker intermolecular forces of attraction than iodine.
- d. Fluorine has stronger intermolecular forces of attraction than iodine.

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RD Boyd



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