CHAPTER 1

MATTER AND MEASUREMENT

Chemistry: deals with composition, structure, and reactions of matter

- Matter has mass and occupies space
 - Mass measures quantity of matter
 - Weight result of gravitational attraction between matter
 - Weight and mass are NOT the same
- Composition what matter is made of and how much of each component is present
 - Expressed by...
 - Weight/mass
 - Volume
 - Percent
 - # of moles
 - # of atoms
 - Macroscopic can be seen and weighed
 - o Submicroscopic level number and types of atoms
 - Atoms = simple units of matter
 - Molecules = combinations of atoms
 - Don't separate when dissolved likes compounds ⇒ ionic compounds don't exist as molecules and separate when dissolved
 - Qualitative = ethanol consists of carbon, hydrogen and oxygen
 - Quantitative = ethanol = C2H6O
- Structure arrangement of components and how they are bonded
 - Different arrangements of same combination of components = isomers
 - o Structures rotate freely
 - Dashed lines = behind

Ethanol

C.H.O

C.H.OH

II II

HH

H-C-C-O-H

- Reactions changes in composition and structure
 - o What products?
 - How much of product? (stoich)
 - How fast change occurs?
 - What energy changes accompany rxn?

exothermic reaction → produces energy over a long period of time

Scientific Method

- Experiment (record observations) → numerical + observational
 - Careful recordings and analysis of data
 - Controlled conditions so we pinpoint what is causing another thing to happen
 - Reproducible
- Draw a Conclusion <u>LAW</u>
 - Concise statement about a basic relationship or regularity of nature drawn from observations

Gradpoint Chemistry Chapter 1

Raymond Chang

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Who's who in New Jersey ,1939 Who's who in Delaware ,1939 Who's who in Maryland ,1939 Chemistry:

 $\underline{\text{chapter 1-11}} \ Edward \ J. \ Neth, Paul \ Flowers, Klaus \ Hellmut \ Theopold, Richard \ Langley, William \ R. \ Robinson, 2017$

Chemistry: Chapters 1-10 Raymond Chang, 2007

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