

Study Card: Gas Variables

Four variables, referred to as **state variables**, describe every sample of gas. They are:

1. Pressure (**P**)
2. Temperature (**T**)
3. Volume (**V**)
4. Number of moles (**n**)

A change in one of these variables will affect the value or one or more of the other variables.

Pressure (P)

Gas pressure is caused by gas particles colliding with the container walls and exerting forces upon the walls.

Pressure = Force / Area

Units:: atm, mm Hg, torr, lb/in², N/m²

$$1 \text{ atm} = 760 \text{ mm Hg} = 14.7 \text{ lb/in}^2$$

Temperature (T)

Temperature reflects the average kinetic energy possessed by gas particles as they move about the container.

High temperature gas samples have particles moving at high speeds. Thus, *a thermometer is like a speedometer.*

Units:: Kelvin (abbreviated K)

Volume (V)

Volume refers to how much space the gas sample occupies.

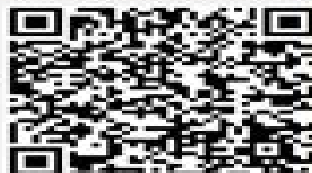
You may know: Gases are expandable and compressible. Changes in T, P, or n often lead to volume changes.

Units: L, mL, m³

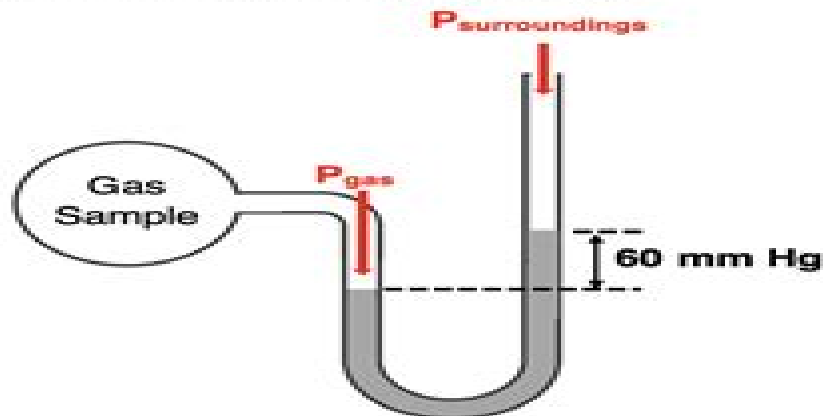
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Manometers are used to measure the pressure of a gas. The gas sample presses down on a column of mercury in a U-tube. The difference in height of the Hg levels can be used to determine the gas pressure.



$$P_{\text{gas}} > P_{\text{surroundings}}$$

60 mm height difference

If $P_{\text{surroundings}} = 740 \text{ mm Hg}$,
Then $P_{\text{gas}} = 800 \text{ mm Hg}$

Number of Moles (n)

Like pair or dozen, the **mole** is a counting unit used to keep count of the large number of small particles in a sample of matter.

$$1 \text{ mole} = 6.022 \times 10^{23} \text{ particles}$$

We express the number of particles of gas in a sample using the mole. The number of moles can be determined by multiplying the mass of the sample in grams by the molar mass of the gas.

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H Kauffman



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