

Gas Laws Worksheet

$$\text{atm} = 760.0 \text{ mm Hg} = 101.3 \text{ kPa} = 760.0 \text{ torr}$$

Boyle's Law Problems:

$$P_1 V_1 = P_2 V_2$$

1. If 22.5 L of nitrogen at 748 mm Hg are compressed to 725 mm Hg at constant temperature. What is the new volume?

$$(748 \text{ mm Hg})(22.5 \text{ L}) = (725 \text{ mm Hg}) V_2$$

$$V_2 = \frac{(748 \text{ mm Hg})(22.5 \text{ L})}{(725 \text{ mm Hg})}$$

$$V_2 = 23.2 \text{ L}$$

2. A gas with a volume of 4.0 L at a pressure of 205 kPa is allowed to expand to a volume of 12.0 L. What is the pressure in the container if the temperature remains constant?

$$(4.0 \text{ L})(205 \text{ kPa}) = (12.0 \text{ L}) P_2$$

$$P_2 = \frac{(4.0 \text{ L})(205 \text{ kPa})}{12.0 \text{ L}}$$

$$P_2 = 68.3 \text{ kPa}$$

3. What pressure is required to compress 196.0 liters of air at 1.00 atmosphere into a cylinder whose volume is 26.0 liters?

$$(196.0 \text{ L})(1.00 \text{ atm}) = (26.0 \text{ L}) P_2$$

$$P_2 = \frac{(196.0 \text{ L})(1.00 \text{ atm})}{26.0 \text{ L}}$$

$$P_2 = 7.54 \text{ atm}$$

4. A 40.0 L tank of ammonia has a pressure of 12.7 kPa. Calculate the volume of the ammonia if its pressure is changed to 8.4 kPa while its temperature remains constant.

$$(40.0 \text{ L})(12.7 \text{ kPa}) = (8.4 \text{ kPa}) V_2$$

$$V_2 = \frac{(40.0 \text{ L})(12.7 \text{ kPa})}{8.4 \text{ kPa}}$$

$$V_2 = 60.5 \text{ L}$$

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Table of Contents Gas Laws Practice Problems With Answers

1. Understanding the eBook Gas Laws Practice Problems With Answers
 - The Rise of Digital Reading Gas Laws Practice Problems With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Laws Practice Problems With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Laws Practice Problems With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Laws Practice Problems With Answers
 - Personalized Recommendations
 - Gas Laws Practice Problems With Answers User Reviews and Ratings
 - Gas Laws Practice Problems With Answers and Bestseller Lists
5. Accessing Gas Laws Practice Problems With Answers Free and Paid eBooks
 - Gas Laws Practice Problems With Answers Public Domain eBooks
 - Gas Laws Practice Problems With Answers eBook Subscription Services
 - Gas Laws Practice Problems With Answers Budget-Friendly Options
6. Navigating Gas Laws Practice Problems With Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Gas Laws Practice Problems With Answers Compatibility with Devices
 - Gas Laws Practice Problems With Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gas Laws Practice Problems With Answers
 - Highlighting and Note-Taking Gas Laws Practice Problems With Answers
 - Interactive Elements Gas Laws Practice Problems With Answers
 8. Staying Engaged with Gas Laws Practice Problems With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Laws Practice Problems With Answers
 9. Balancing eBooks and Physical Books Gas Laws Practice Problems With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Laws Practice Problems With Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Gas Laws Practice Problems With Answers
 - Setting Reading Goals Gas Laws Practice Problems With Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Gas Laws Practice Problems With Answers
 - Fact-Checking eBook Content of Gas Laws Practice Problems With Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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