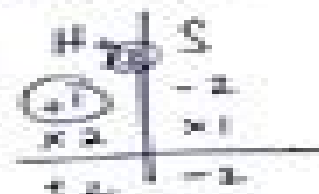
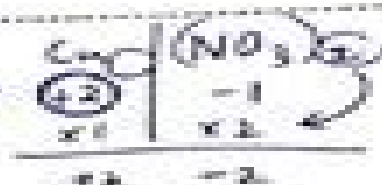


# Worksheet #3: Naming Inorganic Compounds

1. Give the name for each of the following ionic compounds:

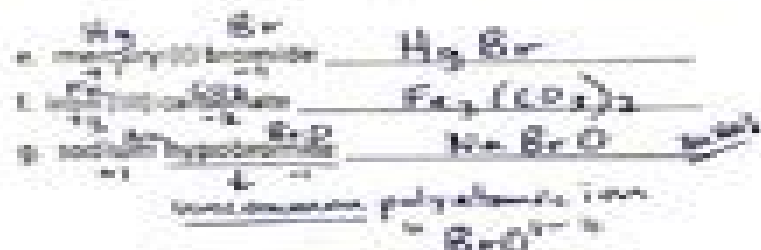
- $\text{AlF}_3$  Aluminum Fluoride
- $\text{Fe(OH)}_2$  Iron(II) Hydroxide
- $\text{Cu(NO}_3)_2$  Copper (II) Nitrate
- $\text{Ba(ClO}_3)_2$  Barium Perchlorate
- $\text{Li}_3\text{PO}_4$  Lithium Phosphate
- $\text{HgS}$  Mercury (I) Sulfide
- $\text{Ca(CH}_3\text{COO)}_2$  Calcium Acetate
- $\text{Cr}_2(\text{CO}_3)_3$  Chromium (III) Carbonate
- $\text{K}_2\text{Cr}_2\text{O}_7$  Potassium Dichromate
- $\text{(NH}_4)_2\text{SO}_4$  Ammonium Sulfate



all of the underlined compounds are polyatomic ions → Memorize Them

2. Write the chemical formula for each of the following compounds:

- Copper(II) oxide  $\text{Cu}_2\text{O}$
- potassium peroxide  $\text{K}_2\text{O}_2$
- aluminum hydroxide  $\text{Al(OH)}_3$
- Iron(III) bromide  $\text{FeBr}_3$



3. Give the name or chemical formula, as appropriate, for each of the following acids:

- $\text{HNO}_3$  nitric acid
- $\text{HBr}$  hydrobromic acid
- $\text{H}_3\text{PO}_4$  phosphoric acid  
phosphate  $\text{PO}_4^{3-}$  (the = oxy)

- hypochlorous acid  $\text{HClO}$
  - iodic acid  $\text{HIO}_3$
  - sulfurous acid  $\text{H}_2\text{SO}_3$
- Band off of Iodate is  $\text{IO}_3^-$

4. Give the name or chemical formula, as appropriate, for each of the following molecular substances:

- $\text{SiF}_4$  silicon tetrafluoride
- $\text{IF}_5$  Iodine pentafluoride
- $\text{XeO}_3$  Xenon trioxide

- dinitrogen tetroxide  $\text{N}_2\text{O}_4$
- hydrogen cyanide  $\text{HCN}$
- tetraphosphorus hexasulfide  $\text{P}_4\text{S}_6$

There are  
all binary  
ionic  
compounds  
(+/-)

3

Binary

3

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3

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**Wolfgang Guggemos**



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