# Hydrolysis -The Reaction of lons with Water

# Lab Report

# Purpose

The purpose of this lab is to predict whether a substance is acidic, basic, or neutral and to use the universal indicator solution to test it.

# **Hypothesis**

Sodium Acetate	$NaCH_3COO_{(a)} \rightarrow Na^*_{(aq)} + CH_3COO^*_{(aq)}$ $CH_3COO^*_{(aq)} + H(OH)_{(b)} \leftrightarrow CH_3COOH_{(aq)} + OH_{(aq)}$		Basic
Sodium Chloride	NaCl <sub>so</sub> ++Na* <sub>seq</sub> + Cl* <sub>seq</sub>		Neutral
Ammonium Chloride	$NH_4Cl_{03}\rightarrow NH_{4(eq)}+Cl_{(eq)}$ $NH_4^*_{(eq)}+H(OH)_{(3}\leftrightarrow NH_4(OH)_{(eq)}+H^*_{(eq)}$		Acidic
Ammonium Sulfate	$(NH_4)_2SO_{4(g)} \rightarrow NH_4^*_{(acg)} + SO_4^{2^*}_{(acg)}$ $NH_4^*_{(acg)} + H(OH)_{(g)} \leftrightarrow NH_4(OH)_{(acg)} + H^*_{(acg)}$ $SO_4^{2^*}_{(acg)} + H(OH)_{(g)} \leftrightarrow HSO_4^*_{(acg)} + OH^*_{(acg)}$	Ka=5,4×10 <sup>-10</sup>	Acidic
		Kb=8.3×10 <sup>-13</sup>	
Aluminum Chloride	$\begin{array}{l} AlCl_{(acj)} \rightarrow Al^{3*}_{(acj)} + Cl^{*}_{(acj)} \\ Al^{3*}_{(acj)} + H(OH)_{(acj)} \rightarrow Al(OH)_{(acj)} + H^{*}_{(acj)} \end{array}$		Acidic
Calcium Nitrate	Ca(NO <sub>3</sub> ) <sub>263</sub> →Ca <sup>2*</sup> <sub>Ceql</sub> + No <sub>3649</sub>		Neutral
Iron(III) Sulfate	$\begin{array}{l} {\sf Fe}_2({\sf SO}_4)_{3(q)} {\to} {\sf Fe}^{3*}{}_{(aq)} + {\sf SO}_4{}^{2*}{}_{(aq)} \\ {\sf Fe}^{3*}{}_{(aq)} + {\sf H}({\sf OH})_{(3} {\leftrightarrow} {\sf Fe}({\sf OH})^{2*}{}_{(aq)} + {\sf H}^*{}_{(aq)} \\ {\sf Fe}^{3*}{}_{(aq)} + {\sf H}({\sf OH})_{(0} {\leftrightarrow} {\sf HFe}^{4*}{}_{(aq)} + {\sf OH}^*{}_{(aq)} \\ {\sf SO}_4{}^{2*}{}_{(aq)} + {\sf H}({\sf OH})_{(3} {\leftrightarrow} {\sf HSO}_4{}^*{}_{(aq)} + {\sf OH}^*{}_{(aq)} \end{array}$	Ka=6.0×10 <sup>-8</sup>	Acidic
		Kb=1.7×10 <sup>-12</sup>	
		Kb=8.3×10 <sup>-13</sup>	
Sodium Carbonate	$Na_2CO_{3(q)} \rightarrow Na^*_{(q)} + CO_{3(q)}$ $CO_{3(q)} + H(OH)_{13} \leftrightarrow HCO_{3(q)} + OH_{(q)}$		Basic

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