

Experiment 10: Electrodes with a Standard Solution

From: *Basic Laboratory Chemistry Experiments*, pp. 45-46

The purpose of a concentration cell is to determine the concentration of an electrolyte. In this experiment, a cell is constructed from two identical half-cells, each containing a standard solution of a metal ion and a metal electrode. The only difference between the two half-cells is the concentration of the metal ion. The cell potential is measured, and the concentration of the metal ion in the unknown half-cell is determined. Other practical considerations, such as the use of salt bridges and the effect of temperature, are also discussed. The experiment is designed to be performed in a laboratory setting.

In this experiment, you will be constructing a concentration cell from two half-cells. Each half-cell consists of a standard solution of a metal ion and a metal electrode. The only difference between the two half-cells is the concentration of the metal ion. The cell potential is measured, and the concentration of the metal ion in the unknown half-cell is determined.

Objectives:

- to construct a concentration cell from two half-cells
- to determine the concentration of the metal ion in the unknown half-cell
- to determine the standard potential of the metal ion
- to determine the standard potential of the metal ion from the cell potential
- to determine the standard potential of the metal ion from the cell potential

Materials:

- standard solution of metal ion
- standard solution of metal ion
- standard solution of metal ion
- standard solution of metal ion
- standard solution of metal ion
- standard solution of metal ion
- standard solution of metal ion
- standard solution of metal ion
- standard solution of metal ion

Procedure:

- Construct a concentration cell from two half-cells.
- Determine the concentration of the metal ion in the unknown half-cell.
- Determine the standard potential of the metal ion.
- Determine the standard potential of the metal ion from the cell potential.
- Determine the standard potential of the metal ion from the cell potential.

Results:

- Calculate the concentration of the metal ion in the unknown half-cell.
- Calculate the standard potential of the metal ion.

Discussion:

- What are your results?
- What are your conclusions?

Heath Chemistry 1lab Experiment 6b Answers

**Brownlee, Fuller, Hancock, Sohon, Whitsi
t, Paul J. Boylan**

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