

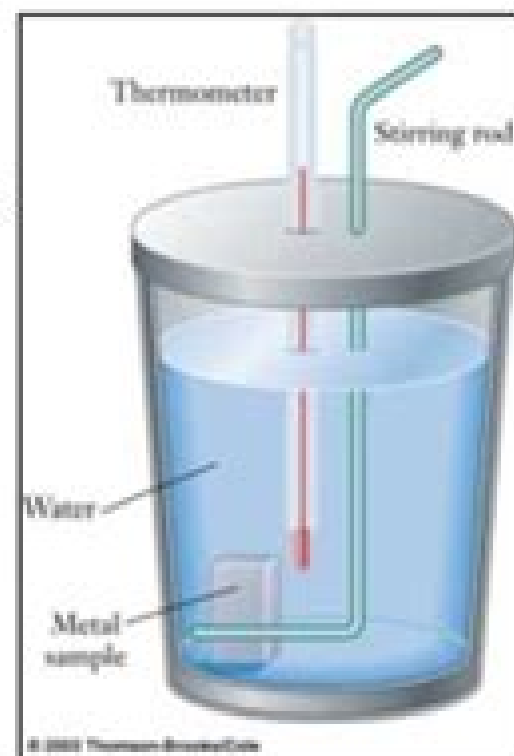
In the laboratory a "*coffee cup*" *calorimeter*, or constant pressure calorimeter, is frequently used to determine the specific heat of a solid, or to measure the energy of a solution phase reaction.

Since the cup itself can absorb energy, a separate experiment is needed to determine the heat capacity of the calorimeter. This is known as *calibrating* the calorimeter and the value determined is called the *calorimeter constant*.

One way to do this is to use a common metal of known heat capacity. In the laboratory a student heats **95.50** grams of **zinc** to **98.11 °C** and then drops it into a cup containing **81.69** grams of water at **23.33 °C**. She measures the final temperature to be **30.61 °C**.

Using the accepted value for the specific heat of **zinc** (See the References tool), calculate the calorimeter constant.

Calorimeter Constant = J/°C.



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Heat Of Neutralization Lab Answer Key

PT Brinkman

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