



Hp Officejet J5750 All In One Manual

Puja Mehta



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 hot object is 60 c the specific heat of copper and water are 0 09 cal gm c and 1 cal gm c respectively

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 temperature change plugging in the values given in the problem you get $q_{\text{water}} = 4.18 \text{ j/g} \cdot \text{c} \times 110 \text{ g} \times 26.6 \text{ c} - 25.0 \text{ c}$ $q_{\text{water}} = 550 \text{ j}$ $\Delta h_{\text{q water}} = 550 \text{ j}$

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 for water 4.184 j/goc 1 what is the specific heat of aluminum if the temperature of a 28.4 g sample of aluminum is increased
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 absorbed at constant pressure is equal to Δh the relationship between heat and Δh_{rxn} is $\Delta h_{\text{rxn}} = q_{\text{rxn}} / \text{mcs} \Delta t$ 7 3
 14 7 3 14 $\Delta h_{\text{rxn}} = q_{\text{rxn}} / \text{mcs} \Delta t$

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web 6 a 218 g sample of steam at 121 °C is cooled to ice at 14 °C find the change in heat content of the system 7 in going from ice at 34 °C to steam at 138 °C a sample of water absorbs 1.41×10^5 J find the mass of the sample answers 1 2 68 x 104 J 2 6 23 x 104 J 3 1 71 x 104 J 4 1 11 x 105 J 5 2 11 x 105 J 6

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