

## Mole Ratios

How can the coefficients in a chemical equation be interpreted?

### Why?

A balanced chemical equation can tell us the number of reactant and product particles (ions, atoms, molecules or formula units) that are necessary to conserve mass during a chemical reaction. Typically when we balance the chemical equation we think in terms of individual particles. However, in real life the reaction represented by an equation occurs an unimaginable number of times. Short of writing very large numbers ( $10^{23}$  or larger) in front of each chemical in the equation, how can we interpret chemical equations so that they more realistically represent what is happening in real life? In this activity you will explore the different ways a chemical reaction can be interpreted.

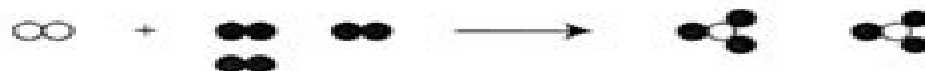
### Model 1 – A Chemical Reaction



1. Consider the reaction in Model 1.
  - a. What are the coefficients for each of the following substances in the reaction?



- b. Draw particle models below to illustrate the reaction in Model 1.



2. Consider each situation below as it relates to the reaction in Model 1.
  - a. Calculate the amount of reactants consumed and products made.
  - b. Record the ratio of  $\text{N}_2$  to  $\text{H}_2$  to  $\text{NH}_3$ . Reduce the ratio to the lowest whole numbers possible.

	$\text{N}_2$ Consumed	$\text{H}_2$ Consumed	$\text{NH}_3$ Produced	Ratio $\text{N}_2:\text{H}_2:\text{NH}_3$ (reduced)
For a single reaction, how many molecules of each substance would be consumed or produced?	1 molecule	3 molecules	2 molecules	1:3:2
If the reaction occurred one hundred times, how many molecules would be consumed or produced?	100 molecules	300 molecules	200 molecules	1:3:2
If the reaction occurred 538 times, how many molecules would be consumed or produced?	538 molecules	1614 molecules	1076 molecules	1:3:2

# Mole Ratios In Chemical Reactions Pogil Answers

**Jane D. Smith, George R. Hague**



## **Mole Ratios In Chemical Reactions Pogil Answers:**

**Mole Ratios** Lifeliqe,2019 This lesson plan relates balanced chemical equations to everyday analogies defines stoichiometry and uses mole ratios to convert between amounts of substances in chemical reactions *Chemistry Equations & Answers* Speedy Publishing,2014-08-01 One of the most important parts of learning chemistry is simply knowing all of the equations and formulas that are used in it It is important to know the exact formatting of these equations and most classes will require a student to know them for exams It is a good idea to learn these equations and formulas with the use of a study pamphlet The pamphlet can condense all of the information so a student can memorize the equations and formulas while studying

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**You Can Do Chemistry** Sunil Tanna,2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018

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Bite Sized Chemistry Calculations is a series of books on chemistry calculations aimed at helping students overcome the challenges associated with tackling the various types of calculations encountered in different aspects of chemistry focusing on a few topics at a time to facilitate comprehension Written by an experienced chemistry educator each book in the series has been tailored to fully meet the needs of students at all levels especially those taking college level general chemistry courses as well as those following various O level curricula worldwide This part of the series explores the different types of problems and calculations encountered in mass the mole and stoichiometry including the determination of formulae of ionic compounds relative formula masses mass and percent compositions of compounds all aspects of mole calculations empirical and molecular formulae calculations based on chemical equations limiting reagents gas stoichiometry and percent yield The series is packed with many salient features that are meant to facilitate both teaching and learning Some of these include helpful explanations many examples alternative ways to solve problems plenty of practice questions complete answers an appendices With this book you will be well prepared for your exams and boost your performance CONTENTS 1 Writing the Formulae of Ionic Compounds 2 Formula Masses 3 Mass and Percent Compositions 4 The Mole and Mass 5 The Mole and Number of Particles 6 The Mole and Concentration 7 The Mole and Molar Volume 8 Empirical and Molecular Formulae 9 Chemical Equations 10 Calculations Based on Chemical Equations 11 Limiting Reagent 12 Gas Volume Volume Stoichiometry 13 Percent Yield Answers to Practice Problems Appendices

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**CliffsNotes Chemistry Practice Pack** Charles Henrickson,2010-02-08 About the Contents Pretest Helps you pinpoint where you need the most help Topic Area Reviews Measurement and Units of Measurement Matter Elements Compounds and Mixtures Atoms I The Basics Formulas and Names of Ionic Compounds Acids and Bases The Mole Elements and Compounds Percent Composition and Empirical and Molecular Formulas Chemical Reactions and Chemical Equations Calculations Using Balanced Equations Atoms II Atomic Structure and Periodic Properties Chemical Bonding The Formation of Compounds Gases and the Gas Laws The Forces between Molecules Solids and Liquids Solutions and Solution Composition Acids Bases and Neutralization Glossary Customized Full Length Exam Covers all subject areas Pretest that pinpoints what you need to study most Clear concise reviews of every topic Targeted example problems in every chapter with solutions and explanations Customized full length exam that adapts to your skill level

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