

CHAPTER 6 REVIEW

The Periodic Law**SECTION 3****SHORT ANSWER** Answer the following questions in the space provided.

1. _____ When an electron is added to a neutral atom, energy is
 - (a) always absorbed.
 - (b) always released.
 - (c) either absorbed or released.
 - (d) neither absorbed nor released.
2. _____ The energy required to remove an electron from a neutral atom is the atom's
 - (a) electron affinity.
 - (b) electron energy.
 - (c) electronegativity.
 - (d) neither absorbed nor released.
3. From left to right across a period on the periodic table,
 - _____ a. electron affinity values tend to become more (negative or positive).
 - _____ b. ionization energy values tend to (increase or decrease).
 - _____ c. atomic radii tend to become (larger or smaller).
4. _____ a. Name the halogen with the least-negative electron affinity.
_____ b. Name the alkali metal with the highest ionization energy.
_____ c. Name the element in Period 3 with the smallest atomic radius.
_____ d. Name the Group 14 element with the largest electronegativity.
5. Write the electron configuration of the following:
 - a. Na

 - b. Na⁺

 - c. O

 - d. O²⁻

 - e. Co²⁺

Modern Chemistry Review Answer Sheet

Puja Mehta



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